

Reg. No. :

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H 2426

B.E./B.Tech. DEGREE EXAMINATION, MAY/JUNE 2009

Second Semester

Mechanical Engineering

ME 132 — THERMODYNAMICS

Time : Three hours

Maximum : 100 marks

(Use of thermodynamic tables permitted)

Answer ALL questions.

PART A — (10 × 2 = 20 marks)

1. Write the difference between open and closed systems.
2. State 'Zeroth law of thermodynamics'.
3. Name the processes in Carnot cycle.
4. What is the relationship between COP of the heat pump and refrigerator?
5. Define 'Mole fraction'.
6. State the laws of perfect gas.
7. What is super saturated steam?
8. Find the specific volume, enthalpy and entropy of saturated steam at 10 bar.
9. What is meant by 'Adiabatic flame temperature'?
10. What is the mass of carbon-dioxide from one kg of carbon combustion?

PART B — (5 × 16 = 80 marks)

11. (a) Determine the heat transfer and its direction for a system in which a perfect gas having molecular weight of 16 is compressed from 101.3 kPa, 20°C to a pressure of 600 kPa following the law $Pv^{1.2} = \text{constant}$. Take specific heat at constant pressure of gas as 1.7 kJ/kgK. (16)

Or

Engineering Thermodynamics Question Paper Pune University

Ensheng Dong



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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1
 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$ one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being NaOH reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{C}_2\text{O}_4 + 2\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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concentration of an unknown substance by mixing 2m hcl with the naoh of unknown concentration in order to experimentally ascertain the concentration of the naoh

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction
 $\text{NaOH (aq)} + \text{C}_6\text{H}_4\text{COOH (aq)} \rightarrow \text{C}_6\text{H}_4\text{COONa (aq)} + \text{H}_2\text{O (l)}$
the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the