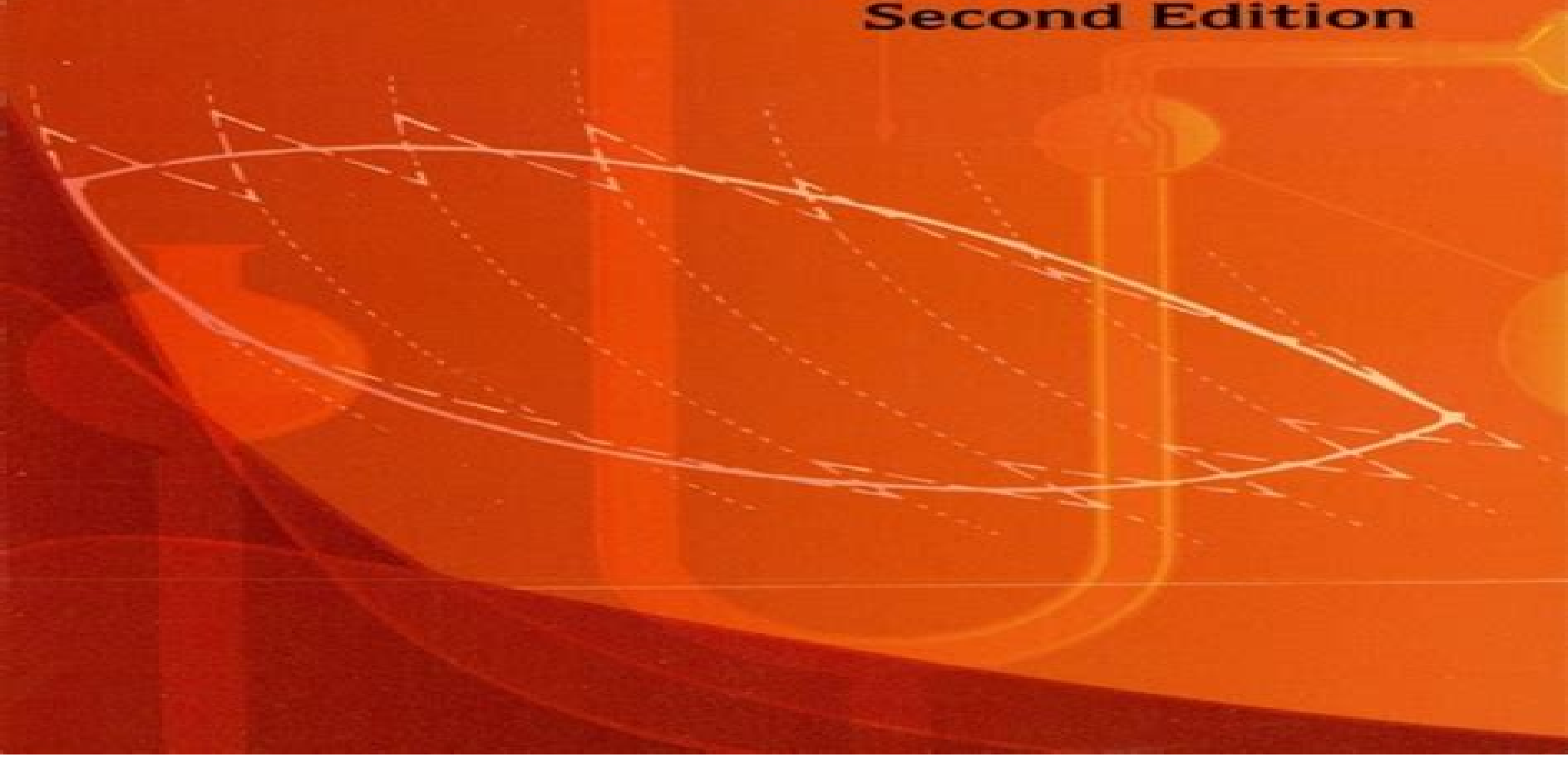


LEONARD K. NASH

Elements of Chemical Thermodynamics

Second Edition



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Leonard K Nash

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you have 0.216232 grams mole 21.3 grams 0.66 moles 32 g m

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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g $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol $\text{Al} \times 1 \text{ mol Al}_2\text{O}_3$ c determine the percent yield

yield actual $\times 100$ $260.0 \text{ g} \times 100 / 86.01$ theoretical 302.3 g

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C is the limiting reagent 4.33 g of H₂ are left over 13

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yield theoretical $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ based on this definition we would expect a percent yield

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web ante acontecimientos como el del 26 de septiembre de 2014 ningún país puede seguir adelante sin conocer la verdad a la que tienen derecho las víctimas y la sociedad los hechos de iguala nos obligan a reflexionar sobre el momento que vive México retratan con crudeza la degradación de las instituciones cuya obligación es procurar