

5

ELECTRONS IN ATOMS

A. Matching

Match each term in Column A with the correct description in Column B. Write the letter of the correct term on the line.

Column A

- _____ 1. the ways in which electrons are arranged around the nuclei of atoms
- _____ 2. the ejection of electrons by metals when light shines on them
- _____ 3. the region around the nucleus of an atom where an electron is likely to be moving
- _____ 4. An atomic orbital may describe at most two electrons.
- _____ 5. the regions within which electrons have the highest probability of being found
- _____ 6. When electrons occupy orbitals of equal energy, one electron enters each orbital until all the orbitals contain one electron.
- _____ 7. the amount of energy required to move an electron from its present energy level to the next highest one
- _____ 8. light quanta
- _____ 9. Electrons enter orbitals of lowest energy first.
- _____ 10. the modern description of the electrons in atoms

Column B

- a. Aufbau principle
- b. quantum
- c. photoelectric effect
- d. atomic orbitals
- e. Pauli exclusion principle
- f. energy level
- g. Hund's rule
- h. electron configurations
- i. quantum mechanical model
- j. photons

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 11. Bohr's contribution to the development of atomic structure:
 - a. was referred to as the "planar parking model."
 - b. was the discovery that electrons surround a dense nucleus.
 - c. was proposed that electrons travel in circular orbits around the nucleus.
 - d. is the quantum mechanical model.
- _____ 12. What is the total number of orbitals in the third principal energy level?
 - a. 1
 - b. 4
 - c. 9
 - d. 16

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Jeffrey C. Pommerville



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