

8 • Electron Configurations & Periodicity

ANSWERS TO STUDY QUESTIONS

- Write the electron configurations of the following elements using the shorthand notation for the noble gas cores.
 - phosphorus [Ne] $3s^2 3p^3$
 - nickel [Ar] $3d^8 4s^2$
 - osmium [Xe] $4f^{14} 5d^6 6s^2$
 - californium [Rn] $5f^{10} 7s^2$
 - titanium [Ar] $3d^2 4s^2$
- Which orbital is filled following these orbitals?
 - 3d is followed by 4p
 - 4s is followed by 3d
 - 5p is followed by 6s
 - 5f is followed by 6d
- How many electrons can be accommodated in
 - a d subshell 10; two in each of 5 orbitals
 - a set of f orbitals 14; two in each of 7 orbitals
 - the $n = 4$ shell 32; $2s^2$
 - the 7s orbital 2
 - a p_x orbital? 2
- What is wrong with the following ground state electron configurations?

<p>a. $\begin{array}{ c c c c } \hline & & & \\ \hline \end{array}$ $\begin{array}{ c } \hline 3d \\ \hline \end{array}$ $\begin{array}{ c c c } \hline & & \\ \hline \end{array}$ $\begin{array}{ c } \hline 4s \\ \hline \end{array}$</p> <p>b. $\begin{array}{ c c c c c } \hline \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow \\ \hline \end{array}$ $\begin{array}{ c } \hline 3d \\ \hline \end{array}$ $\begin{array}{ c } \hline 4s \\ \hline \end{array}$ $\begin{array}{ c c c } \hline \uparrow & \uparrow & \uparrow \\ \hline \end{array}$ $\begin{array}{ c } \hline 4p \\ \hline \end{array}$</p> <p>c. $\begin{array}{ c c c c c } \hline \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow \\ \hline \end{array}$ $\begin{array}{ c } \hline 3d \\ \hline \end{array}$ $\begin{array}{ c } \hline 4s \\ \hline \end{array}$ $\begin{array}{ c c c } \hline \uparrow & \uparrow & \uparrow \\ \hline \end{array}$ $\begin{array}{ c } \hline 4p \\ \hline \end{array}$</p> <p>d. $\begin{array}{ c c c c c } \hline \uparrow & \uparrow & \uparrow & \uparrow & \uparrow \\ \hline \end{array}$ $\begin{array}{ c } \hline 3d \\ \hline \end{array}$ $\begin{array}{ c } \hline 4s \\ \hline \end{array}$ $\begin{array}{ c c c } \hline & & \\ \hline \end{array}$ $\begin{array}{ c } \hline 4p \\ \hline \end{array}$</p> <p>e. $\begin{array}{ c c c c c } \hline \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow & \uparrow\downarrow \\ \hline \end{array}$ $\begin{array}{ c } \hline 3d \\ \hline \end{array}$ $\begin{array}{ c } \hline 4s \\ \hline \end{array}$ $\begin{array}{ c c c } \hline \uparrow\downarrow & & \\ \hline \end{array}$ $\begin{array}{ c } \hline 4p \\ \hline \end{array}$</p>	<p>a. OK, 4s is filled before 3d.</p> <p>b. 4s should be filled before 4p subshell.</p> <p>c. 4s cannot have 2 electrons with same spin.</p> <p>d. OK, stability of half-filled 3d subshell.</p> <p>e. 4p should first fill one orbital at a time.</p>
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- How many unpaired electrons are there in
 - a nitrogen atom 3 unpaired electrons ($2s^2 2p^3$)
 - an iodine atom 1 unpaired electrons ($5s^2 5p^5$)
 - a nickel (II) cation 2 unpaired electrons ($3d^8$)
 - an oxide ion no unpaired electrons (closed shell [Ne])

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Joan I. Senyk, Darrell D. Ebbing



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