

1. How did Bohr expand on Rutherford's model of the atom?

The Rutherford model of the atom was incomplete because it didn't explain how the atom's negatively charged electrons were distributed in the space surrounding the positively charged nucleus.

2. Compare the energy of an electron in the ground state and an electron in the excited state.

A valence electron in the ground state has to absorb energy in order for it to move or jump from the ground state to the excited state. This means that a ground state electron is lower energy than an excited state electron.

3. When an electron falls from a higher level to a lower level, how is the energy released?

When a valence electron absorbs energy in the form of heat or light it uses that energy to jump to an excited state (outer level). An excited (valence) electron then releases that energy in the form of light (and a little heat) and falls back to its ground state. The movement of the valence electron from the ground state to the excited state is momentary.

4. Explain how the gaseous neon atoms in a neon sign emit light.

Gaseous atom (nonmetal atoms) absorb energy in the form of high voltage electrical energy. The valence electrons of the gaseous atoms use this energy to jump from the ground state to an excited state and then emit (give off) that energy in the form of light (and a little heat).

5. List the seven colors of the visible spectrum in order of increasing energy.

Red, Orange, Yellow, Green, Blue, Indigo, and Violet

6. What is the energy difference between a photon of yellow light and a photon of violet light?

A photon of yellow light is lower energy than a photon of violet energy.

7. Determine the type of radiation (gamma rays, infrared waves, or radio waves) that has the:

- a. longest wavelength radio waves
- b. highest frequency gamma rays
- c. greatest energy gamma rays

Electron Configuration Webquest Answer Key

Mark Steyvers

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