

## Practice Problems (Chapter 8): Electron Configuration

CHEM 1A

1. Write complete electron configurations for the following atoms:a) F  $1s^2 2s^2 2p^5$ 

b) Al

c) V

d) Cu

e) Kr

2. Write noble gas electron configurations for the following atoms:a) Cd  $[\text{Kr}] 5s^2 4d^{10}$ 

b) I

c) Ba

d) Rn

3. How many core, outer, and valence electrons do each of the following atoms have?

	Core	Outer	Valence
a) H	0	1	1
b) Si	10	4	4
c) Ti	18	2	4
d) Br	28	7	7
e) Ag	36	1	11

(3d  $e^-$  are core  $e^-$  for rep. elements)

(an exception)

4. Write orbital notations for the following atoms:

a) O  $\frac{1\downarrow}{1s} \frac{1\downarrow}{2s} \frac{1\downarrow}{2p} \frac{1}{2p} \frac{1}{2p}$   $\frac{\quad}{3s}$   $\frac{\quad}{3p} \frac{\quad}{3p} \frac{\quad}{3p}$   $\frac{\quad}{4s}$   $\frac{\quad}{3d} \frac{\quad}{3d} \frac{\quad}{3d}$ b) P  $\frac{1\downarrow}{1s} \frac{1\downarrow}{2s} \frac{1\downarrow}{2p} \frac{1\downarrow}{2p} \frac{1\downarrow}{2p}$   $\frac{1\downarrow}{3s}$   $\frac{1}{3p} \frac{1}{3p} \frac{1}{3p}$   $\frac{\quad}{4s}$   $\frac{\quad}{3d} \frac{\quad}{3d} \frac{\quad}{3d}$ c) Cr (an exception)  $\frac{1\downarrow}{1s} \frac{1\downarrow}{2s} \frac{1\downarrow}{2p} \frac{1\downarrow}{2p} \frac{1\downarrow}{2p}$   $\frac{1\downarrow}{3s}$   $\frac{1\downarrow}{3p} \frac{1\downarrow}{3p} \frac{1\downarrow}{3p}$   $\frac{1}{4s}$   $\frac{1}{3d} \frac{1}{3d} \frac{1}{3d} \frac{1}{3d} \frac{1}{3d}$ d) Co  $\frac{1\downarrow}{1s} \frac{1\downarrow}{2s} \frac{1\downarrow}{2p} \frac{1\downarrow}{2p} \frac{1\downarrow}{2p}$   $\frac{1\downarrow}{3s}$   $\frac{1\downarrow}{3p} \frac{1\downarrow}{3p} \frac{1\downarrow}{3p}$   $\frac{1\downarrow}{4s}$   $\frac{1\downarrow}{3d} \frac{1\downarrow}{3d} \frac{1}{3d} \frac{1}{3d} \frac{1}{3d}$

# Electron Configuration Problems

**Julia Burdge**



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