

- Q.20 For a reversible reaction forward and backward rate constant respectively are 2.5×10^{-4} and 7.5×10^{-4} . What is the value of equilibrium constant [RPMT-2001]
 (1) 3 (2) $1/3$ (3) 1 (4) None
- Q.21 If the equilibrium constant for the reaction $N_2 + 3H_2 \rightarrow 2NH_3$ is K then the equilibrium constant for the reaction $2N_2 + 6H_2 \rightarrow 4NH_3$ would be equal to [RPMT-2002]
 (1) K^2 (2) $\sqrt{K}$ (3) $\frac{1}{\sqrt{K}}$ (4) $\frac{1}{K^2}$
- Q.22 The reaction quotient (Q) for the reaction $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$ is given by $Q = \frac{[NH_3]^2}{[N_2][H_2]^3}$. The reaction will proceed from right to left if [CBSE-PMT-2003]
 (1) $Q < K_e$ (2) $Q > K_e$ (3) $Q = 0$ (4) $Q = K_e$
- Q.23 The following equilibria are given [CBSE-PMT-2003]
 $N_2 + 3H_2 \rightarrow 2NH_3$ K_1
 $N_2 + O_2 \rightarrow 2NO$ K_2
 $H_2 + \frac{1}{2}O_2 \rightarrow H_2O$ K_3
 The equilibrium constant of the reaction $2NH_3 + \frac{5}{2}O_2 \rightarrow 2NO + 3H_2O$, in terms of K_1 , K_2 and K_3 is :
 (1) $K_1 K_2 / K_3$ (2) $K_1 K_3^3 / K_2$ (3) $K_2 K_3^3 / K_1$ (4) $K_1 K_2 K_3$
- Q.24 For the gas phase reaction,
 $C_2H_4 + H_2 \rightleftharpoons C_2H_6$ ($\Delta H = -32.7$ kcal),
 carried out in a vessel, the equilibrium conc. of C_2H_4 can be increased by [AMU 2003]
 (1) Decreasing the pressure (2) Increasing the temperature
 (3) Removing some C_2H_6 (4) Adding some H_2
- Q.25 Consider the following reversible reaction at equilibrium, $2H_2O(g) \rightleftharpoons 2H_2(g) + O_2(g)$
 $\Delta H = 24.7$ kJ. Which one of the following changes in conditions will lead to maximum decomposition of $H_2O(g)$? [Kerala CEE 2004]
 (1) Increasing both temperature and pressure
 (2) Decreasing temperature and increasing pressure
 (3) Increasing temperature and decreasing pressure
 (4) Increasing temperature at constant pressure
- Q.26 In the reaction, $A_2(g) + 4B_2(g) \rightleftharpoons 2AB_4(g)$ $\Delta H < 0$, the formation of $AB_4(g)$ will be favoured at [MP PMT 2004]
 (1) High temperature, low pressure (2) Low temperature, high pressure
 (3) High temperature, high pressure (4) Low temperature, low pressure

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Brendan G. Carr



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