

16. Which of the following substances, when added to pure water, SIGNIFICANTLY increases the water's conductivity?

- (A) Graphite
- (B) Iron(II) hydroxide
- (C) Copper(II) hydroxide
- (D) Sodium chloride

17. A solution having a pH of 1 will

- (A) neutralize a solution of pH 4
- (B) react with zinc metal to produce hydrogen
- (C) react with zinc metal to produce a solution of pH 10
- (D) react with hydrochloric acid to produce a salt and water

18. Which of the following halogens is a liquid at room temperature?

- (A) Chlorine
- (B) Bromine
- (C) Fluorine
- (D) Iodine

19. A "weak acid" is BEST described as one that yields a

- (A) low reading on the pH meter
- (B) low concentration of acid in aqueous solution
- (C) high concentration of hydrogen ions in solution
- (D) low concentration of hydrogen ions in aqueous solution

20. Which of the following methods may be used in the preparation of barium sulfate in the laboratory?

- (A) Distillation
- (B) Evaporation
- (C) Precipitation
- (D) Crystallization

21. Which TWO of the following statements describe acid salts?

- I. Their anions cannot dissociate to give hydrogen ions.
- II. They form a normal salt when all replaceable hydrogen of the acid is removed.
- III. Their anions can dissociate to give hydrogen ions.
- IV. They are not capable of generating a normal salt.

- (A) I and III
- (B) I and IV
- (C) II and III
- (D) III and IV

22. In which of the following reactions is sulfur dioxide acting as an oxidizing agent?

- (A)  $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_3(\text{g})$
- (B)  $\text{SO}_2(\text{g}) + 2\text{H}_2\text{S}(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + 3\text{S}(\text{s})$
- (C)  $\text{SO}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{SO}_3(\text{aq})$
- (D)  $\text{SO}_2(\text{g}) + 2\text{HNO}_3(\text{aq}) \rightarrow \text{H}_2\text{SO}_4(\text{aq}) + 2\text{NO}_2(\text{g})$

23. In which of the following compounds does hydrogen have a negative oxidation number?

- (A)  $\text{CH}_4$
- (B)  $\text{H}_2\text{O}_2$
- (C)  $\text{NH}_3$
- (D)  $\text{NaH}$

# June 2ocr Chemistry Past Paper F332

**J Spring**



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