

Question 5

1. A chemical reaction involving a weak acid and its conjugate base is shown below. The equilibrium constant for this reaction is $K_a = 4.6 \times 10^{-5}$. Calculate the pH of a 0.10 M solution of the weak acid.

$$K_a = \frac{[H^+][A^-]}{[HA]}$$

2. Calculate the pH of a 0.10 M solution of the weak acid. Give your answer to two decimal places.

$$0.1 = \sqrt{(4.6 \times 10^{-5})(0.1)} = 0.00214$$

$$-\log(0.00214) = 2.67$$

$$pH = 2.67$$

3. The chemical reaction below is a buffer reaction. Calculate the pH of a buffer solution, with 0.10 M of 0.10 M acetic acid and 0.10 M of 0.10 M acetate ion.

4. Calculate the pH of the buffer solution.

Give your answer to two decimal places.

$$0.1 \times \frac{1}{10} = 0.01$$

$$0.01 \times \frac{1}{10} = 0.001$$

$$(4.6 \times 10^{-5}) \times \frac{0.01}{0.001} = 0.00046$$

$$-\log(0.00046) = 3.34$$

$$3.34$$

5. A chemical reaction involving a weak acid and its conjugate base is shown below. The equilibrium constant for this reaction is $K_a = 4.6 \times 10^{-5}$. Calculate the pH of a 0.10 M solution of the weak acid.

6. Calculate the pH of a 0.10 M solution of the weak acid. Give your answer to two decimal places.

$$0.1 = \sqrt{(4.6 \times 10^{-5})(0.1)} = 0.00214$$

$$-\log(0.00214) = 2.67$$

7. Calculate the pH of a 0.10 M solution of the weak acid. Give your answer to two decimal places.

$$0.1 = \sqrt{(4.6 \times 10^{-5})(0.1)} = 0.00214$$

8. Calculate the pH of a 0.10 M solution of the weak acid. Give your answer to two decimal places.

June 2014 Chemistry Ocr Unofficial Mark Scheme

Raffaela Di Napoli



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