

Question Number	Acceptable Answer	Reject	Mark
21 (a) (vi)	Colourless / pale yellow to (pale) pink / first permanent pink	Purple Just '(pale) pink'	1

Question Number	Acceptable Answer	Reject	Mark
21 (a) (vii)	<p>(More manganate(VII) is needed to oxidize <math>\text{Fe}^{2+}</math>, so) titre will be larger (1)</p> <p>Stand alone mark</p> <p>Because the Mn oxidation number changes from 7 to 4 (rather than 2) OR Mn accepts fewer electrons per mole (1)</p> <p>(Brown precipitate is) manganese(IV) oxide / <math>\text{MnO}_2</math> ALLOW <math>\text{Mn}(\text{OH})_4</math> (1)</p> <p>IGNORE References to inaccurate / inconsistent titre values</p>	$\text{Mn}(\text{OH})_2$	3

Question Number	Acceptable Answer	Reject	Mark
21 (b) (i)	<p>Anodic area: <math>\text{Fe}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Fe}</math> (<math>E^{\circ} = -0.44 \text{ V}</math>) OR <math>\text{Fe} \rightleftharpoons \text{Fe}^{2+} + 2\text{e}^{-}</math> (1)</p> <p>Cathodic area: <math>\text{O}_2 + 2\text{H}_2\text{O} + 4\text{e}^{-} \rightleftharpoons 4\text{OH}^{-}</math> (<math>E^{\circ} = +0.40 \text{ V}</math>) (1) ALLOW <math>\frac{1}{2}\text{O}_2 + 2\text{H}^{+} + 2\text{e}^{-} \rightleftharpoons \text{H}_2\text{O}</math> (<math>E^{\circ} = +1.23 \text{ V}</math>)</p> <p>Penalise omission of electrons or use of cell diagrams once only</p> <p>Anode and cathode reversed max 1.</p> <p>IGNORE State symbols even if incorrect Single arrow in equations</p>		2

# June 2015 Ial Mark Scheme Chemistry Unit

**Padhraic Smyth**



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