

Lab 12 A - Investigating Chemical Equilibrium

Purpose: The purpose of this lab is to observe and record the shifts in equilibrium caused by changes in concentration and temperature.

Chem Reactions:

Part 1: Equilibrium involving Chromium(VI) Ions:

Table 1

| Reagent Added | Reagent | Color change and no. of drops required | Direction of Equilibrium shift |
|---------------|---------|--|--------------------------------|
| NaCl (step 1) | no | no change | no shift |
| NaOH (step 2) | no | yellow (2) | products |
| HCl (step 3) | yes | green (20) | reactants |
| NaOH (step 4) | yes | blue (24) | reactants |

Part 2: Equilibrium involving Cobalt (II) Ions:

Table 2

| Reagent Added | Reagent | Color Change | Direction of Equilibrium shift |
|----------------------|---------|----------------|--------------------------------|
| Step 1 (water added) | yes | pink | reactants |
| Step 2 (NaOH) | yes | blue/purple | products |
| Step 3 (NaOH) | yes | light red/pink | reactants |

Part 3: Equilibrium involving Chromate and Dichromate ions:

Table 3

| Step 1: 0.1 M and 0.1 M | 0.1 M K ₂ CrO ₄ | 0.1 M H ₂ SO ₄ | 0.1 M K ₂ CrO ₄ |
|-------------------------|---------------------------------------|--------------------------------------|---------------------------------------|
| | yellow | orange | orange |
| | 0.1 M NaOH added | 0.1 M NaOH added | 0.1 M NaOH added |
| | green | green | green |
| Step 2 | 0.1 M K ₂ CrO ₄ | 0.1 M H ₂ SO ₄ | 0.1 M K ₂ CrO ₄ |

Investigating Chemical Equilibrium Lab Report

University of Michigan

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