Part 2: Le Chatelier's Principle: Predicting Shifts in Equilibrium

Le Chatelier's Principle:

When a system at equilibrium is stressed, so that it is no longer is at equilibrium, the reaction system shifts in the forward or reverse direction to reestablish equilibrium.

Stresses include:

- Adding or removing a reactant or product from the system
- Changing the temperature (which changed K_{eq} value)

2. Consider the following chemical equilibrium that is endothermic as written:

Heat +
$$Co(H_2O)_6^{2+}(aq)$$
 + $4Cl^-(aq) \Rightarrow [CoCl_4]^{2-}(aq)$ + $6H_2O(I)$
Pink Blue

At equilibrium at a particular temperature, the concentration of $Co(H_2O)_6^{2+}$ is slightly greater than the concentration of $[CoCl_4]^{2-}$. The color of the solution is therefore a light purple.

Predict what will happen when NaCl is added to the solution. Remember that NaCl(aq) consists of $Na^*(aq)$ and $Cl^*(aq)$ ions.

- (a) The reaction will shift to the products
- (b) Concentration of Co(H₂O)_e²⁺(sq) will decrease
- (c) Concentration of [CoCl₄]²⁻(aq) will _____increase____
- (d) The color of the solution will become more ____blue__
- 3. Consider again the following chemical equilibrium that is endothermic as written:

Heat +
$$Co(H_2O)s^{2^+}(aq)$$
 + $4C\Gamma(aq) \Rightarrow [CoCls]^{2^-}(aq)$ + $6H_2O(I)$

Predict what will happen when the temperature of the equilibrium system is reduced. This is equivalent to removing "heat" from the system.

- (a) The reaction will shift to the reactants
- (b) Concentration of Co(H₂O)₆²⁺(aq) will ____increase____
- (c) Concentration of [CoCl₄]² (aq) will ______decrease_____
- (d) The color of the solution will become more ____pink____
- (c) Does the value of K_{eq} increase or decrease? Remember that K_{eq} is a ratio of products over reactants at equilibrium.

at equilibrium the constant will remain the same

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