

Appendix D

ANSWERS TO SOME PROBLEMS

1-1. (a) At standard temperature and pressure, a mole of an ideal gas contains 6.022×10^{23} molecules (Avogadro's number) and occupies 22.4 liters. Hence, the number per m^3 is $6.022 \times 10^{23}/2.24 \times 10^{-2} = 2.69 \times 10^{25} \text{ m}^{-3}$.

(b) Since $PV = NRT$, $n = N/V = P/RT$. Hence $n_1/n_2 = P_1T_2/P_2T_1$. Taking n_2 to be the density in part (a) and n_1 to be that in part (b), we have

$$n_1 = (2.69 \times 10^{25}) \frac{10^{-3}}{760} \frac{273}{(273 + 20)} = 3.30 \times 10^{19} \text{ m}^{-3}$$

Note that a diatomic gas such as H_2 will have twice as many atoms per torr as, say, He.

1-2. Consider the integral

$$I^2 = \int_{-\infty}^{\infty} e^{-x^2} dx \int_{-\infty}^{\infty} e^{-y^2} dy = \iint_{-\infty}^{\infty} e^{-(x^2+y^2)} dx dy$$

in a two-dimensional space. Transforming to cylindrical coordinates, we have

$$\begin{aligned} I^2 &= \iint e^{-r^2} r dr d\phi = 2\pi \int_0^{\infty} e^{-r^2} r dr \\ &= \pi \int_0^{\infty} e^{-r^2} d(r^2) = -\pi e^{-r^2} \Big|_0^{\infty} = \pi \end{aligned}$$

Hence,

$$I = \int_{-\infty}^{\infty} e^{-x^2} dx = \sqrt{\pi}$$

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Rachel S Tattersall



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