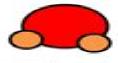
Chapter 6: Chemical Bonding Guided Notes

- B. Why do Atoms Combine?
 - A. Chemical Formula
 - The chemical formula shows:
 - a. The _______ in the compound.
 - b. The ratio of their

← H₂O---



- B. Chemical Bond
 - Strong attractive force between ______ or ions in a molecule or compound.
 - 2. Formed by:
 - electrons, (e-) by losing or gaining electrons.
 - electrons (e-)
- C. Stability
 - 1. Octet Rule
 - a. Most atoms form bonds in order to have _______ valence electrons (e-).
 - b. outer energy level.
 - c. Want to be like the Noble Gases (group 18).
 - Stability is the driving force behind bond formation. All atoms bond to become stable.
 - 3. Transferring e-



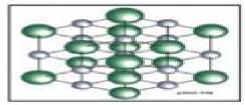
4. Sharing e-



- at. Kinds of Chemical Bonds
 - A. Jonic Bond
 - Attraction between 2 _____ charged ions.



- 2. Ions are ______atoms.
- A ________ is a ______ charged ion. Forms when an atom loses an electron. Formed by
- 4. A ______ is a _____ charged ion. Forms when an atom gains an electron. Formed by ______.
- Formed by electrons (e-) from a to a
- lons form a 3-D crystal lattice.



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Annelies Wilder-Smith

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