Relative Atomic Mass

Definition:

The relative atomic mass, A_r , of an element is defined as the average mass of the isotopes an atom of the element compared with $\frac{1}{12}$ of the mass of an atom of carbon-12.

Note: * Symbol for the relative atomic mass is As and As has no unit.

The values of the Ar of atoms of the elements are given in the Periodic Table and is written on the top of element in the periodic table.

Atomic Number

 It is the number of protons present in the nucleus of an atom. It is also known as the Proton Number.

Mass Number/ Nucleon Number

It is the sum of protons and neutrons present in the nucleus of an atom.

Relationship between Nucleon number and Relative atomic mass, Ar

- The mass of an atom is due mainly to the total number of protons and neutrons in its nucleus (Nucleon number). Thus, we would expect the A_t of the element to be the same as its nucleon number. Thus is one for elements which have no isotopes.
- For most elements, the nucleon number and the A_r are not the same due to the existence of isotopes
- The nucleon number is always a whole number (since we can't have fractions of protons or neutrons in the nucleus), but the A_t values may not be whole numbers. If the Ar is not a whole number and is different from nucleon number, we can conclude that the element has naturally occurring isotopes.

Relative Molecular Mass

- Many elements and compounds exist as molecules.
- The mass of a molecule is measured in terms of its relative molecular mass, Mr

Definition: The relative molecular mass, M_r , of a substance is defined as the average mass of a molecule of the substance compared with 12 to f the mass of Carbon -12 atom

- Relative molecular mass is often abbreviated to molecular mass.
- Symbol is M_t and M_t has no units.
- It is calculated as the sum of the atomic masses of all the atoms in the formula.

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