

Ideal Gas Law Practice Worksheet

Solve the following problems using the ideal gas law:

 How many moles of gas does it take to occupy 120 liters at a pressure of 2.3 atmospheres and a temperature of 340 K?

V= 120L P= 2.3atm T= 340 K n= ?

2) If I have a 50 liter container that holds 45 moles of gas at a temperature of 200°. C, what is the pressure inside the container?

V= 50. L n=45mo1 T=200,°C=473K P=?

3) It is not safe to put aerosol canisters in a campfire, because the pressure inside the canisters gets very high and they can explode. If I have a 1.0 liter canister that holds 2 moles of gas, and the campfire temperature is 1400°, C, what is the pressure inside the canister?

V=1.0 L n=2mol T=1400°C=1673K P= ?

4) How many moles of gas are in a 30 liter scuba canister if the temperature of the canister is 300 K and the pressure is 200 atmospheres?
V= 30. L
(289. a+m \(\) 30. L
(289. a+m \(\) 30. L
) = n \((0.082 \) (0.082 \) (0.082 \)

V= 30. L T= 300. K P= 200. a+m n= ?

5) I have a balloon that can hold 100 liters of air. If I blow up this balloon with 5.0 moles of oxygen gas at a pressure of atmosphere, what is the temperature of the balloon?

> V= 100, L n=3mol P= latm T= ?

Ideal Gas Law Practice Work Sheet

RJ Alexander

Ideal Gas Law Practice Work Sheet:

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