

Graphing Exercise #1

Richard C. Weast, ed., "Data for Plots" (Red gas) Equations
Chemical Data, 1969, McGraw-Hill

Equation numbers from:

$$E^0 = \alpha RT$$

For this exercise for the subject and student, we have a few assumptions about the behavior of the gas. We will use the van der Waals equation of state (Eq. 22.303).

1. The gas is a diatomic molecule.
2. The gas is a diatomic molecule and is a gas at the temperature of the experiment.
3. The gas is a diatomic molecule.

Using the data in the table, we will calculate the values of the parameters α and β for the van der Waals equation of state. The van der Waals equation of state is given by:

$$\left[P + \left(\frac{a}{V^2} \right) \right] (V - b) = \alpha RT$$

The two terms that are indicated are α and β with units of $\text{J} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}$ and $\text{J} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}$.

For a diatomic molecule, the van der Waals equation of state is given by:

$\alpha = \frac{2}{3} \left(\frac{a}{V} \right) \left(\frac{b}{V} \right) \left(\frac{1}{T} \right)$ and $\beta = \frac{1}{3} \left(\frac{a}{V} \right) \left(\frac{b}{V} \right) \left(\frac{1}{T} \right)$

For a diatomic molecule, the van der Waals equation of state is given by:

Ideal Gas Graphing Exercise Answers

KJ Lindholm-Leary

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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