

## Chapter 5 (The Periodic Table) Notes

- Sec 5.6 (Blocks and Sublevels)
  - We can use the periodic table to predict which sublevel is being filled by a particular element.
    - Refer to the periodic table and state the highest energy sublevel for each of the following elements:
      - He – 1s
      - K – 1s
      - U – 5f
      - Pd – 4d
      - Be – 2s
      - Co – 3d
      - Si – 3p
      - Pt – 5d
  - The electron configuration for Na is:  $1s^2 2s^2 2p^6 3s^1$
  - We can abbreviate the electron configuration by indicating the innermost electrons with the symbol of the preceding noble gas.
  - The preceding noble gas with an atomic number less than sodium is neon, Ne.
- Sec 5.7 (Valence Electrons)
  - When an atom undergoes a chemical reaction, only the outermost electrons are involved.
  - These electrons are of the highest energy and are furthest away from the nucleus. These are the valence electrons.
  - The roman numeral in the American convention indicates the number of the valence electrons.
    - Group IA elements have one valence electron.
    - Group VA elements have five valence electrons.
  - When using the IUPAC designations for group numbers the last digit indicates the number of valence electrons.
    - Group 14 elements have four valence electrons.
    - Group 2 elements have two valence electrons.
- Sec 5.8 (Electron Dot Formulas)
  - An **electron dot formula** of an element shows that symbol of the element surrounded by its valence electrons.
    - We use one dot for each valence electron.

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