

Atoms and Bonds Study Guide

Section 1 Fill in the Table

Compound/Molecule	Metal + Nonmetal or Nonmetal + Nonmetal?	Ionic or Covalent Bond?	e ⁻ shared or transferred?
NaCl			
C ₂ H ₆			
PF ₅			
MgO			

Section 2 Periodic Symbol Analysis

74
W
183.84

Element name _____
 Name English or Latin in origin? _____
 Atomic number _____ Number of protons _____
 Atomic mass _____ Number of electrons _____
 Number of neutrons (Mass - # protons) _____
 Metal, nonmetal, or noble gas? _____

Section 3 Properties of Bonds Write "I" if the property describes an ionic bond, "C" if the property describes a covalent bond, and "B" if the property describes both kinds of bonds.

_____ High melting points _____ Weaker bond _____ Electrons shared
 _____ Conducts electricity in water _____ Helps most atoms achieve the octet of a noble gas
 _____ Polar or nonpolar molecules _____ Attraction of opposite charges
 _____ Poor conductors of electricity _____ Crystal shape _____ Low melting points
 _____ Attraction between a cation and an anion _____ Formed or broken in a chemical reaction

Section 4 Formula Writing Write the formula for each of the compounds below. Be sure to balance the charges with subscripts. [Hint: Parenthesized Roman numerals are the charges of transition metals]

- Copper (II) sulfate _____
- Mercury (I) oxide _____
- Mercury (II) nitrate _____
- Beryllium nitrate _____
- Aluminum oxide _____
- Magnesium hydroxide _____

Sulfate = SO₄²⁻Oxide = O²⁻Nitrate = NO₃⁻Hydroxide = OH⁻

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Harvard Project Physics



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