

- Directions: (1) Substantiate your answers with formulas for finding using one or more of the following:
 (2) Balance each equation.
 (3) Identify the type chemical equation, decomposition, single replacement, double replacement, or combination.

1. Aluminum metal reacts with oxygen (air) to form aluminum oxide and oxygen diatomic molecules.

Type: single replacement

Reaction: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$

2. Iron(III)oxide reacts with carbon (C) to form

Type: single replacement

Reaction: $\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 2\text{Fe} + 3\text{CO}$

3. Calciumphosphorus pentoxide can be produced from the elements calcium and phosphorus. Write the balanced chemical equation for this reaction. It is combination.

Type: combination

Reaction: $5\text{Ca} + 2\text{P}_2\text{O}_5 \rightarrow \text{Ca}_5\text{P}_2\text{O}_{10}$

4. Ammonia(gas) reacts to produce nitrogen gas and hydrogen(gas) and water

Type: decomposition

Reaction: $2\text{NH}_3 \rightarrow \text{N}_2 + 3\text{H}_2$

5. Ammonia(gas) reacts with oxygen gas to produce water (l) and nitrogen gas.

Type: decomposition

Reaction: $4\text{NH}_3 + 3\text{O}_2 \rightarrow 2\text{N}_2 + 6\text{H}_2\text{O}$

Even More Fun With Equations 6 33

Charles Hutton



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