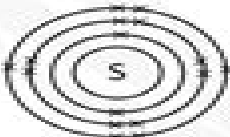
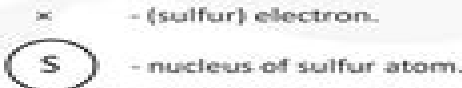
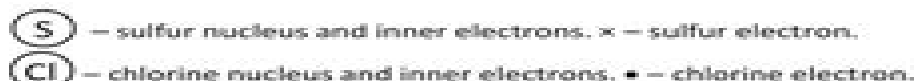


Question 1

- 1(b)(i) Reaction rate decreased as time increased.
- 1(b)(ii) The smaller zinc particles gave a larger surface area which led to a faster reaction rate since there were more successful collisions per unit time.
- 1(b)(iii) Temperature, pressure (for gases) and concentration (of reagents in solution).
- 1(c) One safety precaution could have been the wearing of safety glasses/goggles to protect the eyes from possible acid spray.
- 1(d) Temperature could have been controlled during the experiment by the use of a water bath.
- 1(e)(i) $\text{Zn}_{(s)} + \text{H}_2\text{SO}_{4(aq)} \rightarrow \text{ZnSO}_{4(aq)} + \text{H}_{2(g)}$
- 1(e)(ii) From the balanced equation for the reaction, 1 mole of $\text{Zn}_{(s)}$ produces 1 mole of $\text{H}_{2(g)}$.
 Moles of zinc used = mass of Zn used/molar mass of Zn = $(0.55/65.4)$ mol, i.e. 0.0084 mol. Zn.
 Therefore 0.0084 moles of $\text{H}_{2(g)}$ were produced.
 Volume of $\text{H}_{2(g)}$ produced = moles of $\text{H}_{2(g)}$ produced $\times$ molar volume of gas at RTP.
 Volume of $\text{H}_{2(g)}$ produced = $(0.0084 \times 24\,000) \text{ cm}^3 = 201.8 \text{ cm}^3$.

Question 2

- 2(a) The number of occupied electron shells in an tells us what period it is in. All these elements have three occupied electron shells and are therefore in period three.
- 2(b) Element: Sodium (Na).
 Equation: $\text{Na}_2\text{O}_{(s)} + \text{H}_2\text{O}_{(l)} \rightarrow 2\text{NaOH}_{(aq)}$
- 2(c) Allotrope – one of multiple (2 or more) different forms/structures of the same element, which exist in the same physical state.
- 2(d)  
- 2(e)(i) When sulfur reacts with chlorine there is covalent bonding.
- 2(e)(ii) Both sulfur and chlorine are non-metals. It is energetically preferable for these atoms to share electrons with each other in order to achieve stable outer electronic configurations of 8 electrons, rather than to lose or gain electrons.
- 2(f) SCl_2 .

- 2(g)  

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