

- valve and valve seats for wear or damage
- head gasket

Crankcase Breasts (minimum)

Water Movement ..... 35.4 cm (14 in.)

Hg (mercury) Movement ..... 18.7 mm (3/4 in.)

## Cylinder Leakdown Test

Reason:

To check sealing of piston rings and cylinder bore

Equipment:

- JTD0002 Cylinder Leak Tester

Procedure:

1. Remove head assembly (See "Head Removal and Installation" on page 342 in the Miscellaneous section)



**CAUTION:** Avoid injury! Engine will be HOT. Do not touch with bare skin, especially the exhaust pipe or muffler while making test.

2. Warm engine to operating temperature, then stop engine.
3. Move throttle and lever to HCU/TBAC. Lock park brake.



4. Remove the flywheel screen. Use a socket with a breaker bar (A) on the flywheel nut to hold the flywheel in position when performing the test.



**CAUTION:** Avoid injury! Before attaching adapter hose (B) hose to tester, make sure another person is holding on tight to the breaker bar and the engine is at TDC. When compressed air is applied, there will be considerable torque forcing piston and crankshaft to turn.



5. Connect adapter hose (B) to tester. Warm and record gauge reading while listening for air leaks at carburetor inlet and exhaust outlet. Pressure should hold steady or reduce only slightly.
6. Remove test equipment and install spark plug. Tighten spark plug to 20 ft-lbs (26 Nm.)
7. Install head.

Results:

**NOTE:** A small amount of air escaping from the crankcase breather is normal.

- Excessive air escaping from the crankcase breather indicates worn piston rings or cylinder wall, or a leaking head gasket.
- Air escaping from the carburetor indicates a worn intake valve or seat.
- Air escaping from the exhaust pipe indicates a worn exhaust valve or seat.
- Gauge reading in the Green (low) area indicates good

compression. (Less than 20% is considered normal.)

- Gauge reading in the Yellow (moderate) area indicates borderline compression. The engine is still usable but an overhaul should be considered.
- Gauge reading in the Red (high) area indicates excessive compression loss. Engine overhaul is necessary.

## Fuel and Air Repair

### Upper Blower Housing Removal and Installation

Procedure:



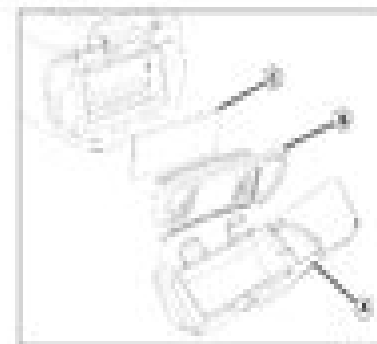
1. Remove four screws (A) holding the outer blower housing in place.
2. Lift outer blower housing up and off of engine.
3. Remove the screws holding the dipstick tube in place.
4. Install in reverse order of removal.

Specification:

Blower Housing Screws ..... 10 ft-lbs (13 Nm.)

## Air Cleaner Service

Procedure:



**IMPORTANT:** Avoid damage! DO NOT use petroleum solvents, such as kerosene, to clean paper cartridge. They may cause cartridge to deteriorate. DO NOT oil paper cartridge. DO NOT use pressurized air to clean or dry paper cartridge.

1. Remove cover (A) to service dust element or cleaner.
2. Remove cartridge (B) and foam pre-cleaner (C).
3. Wash pre-cleaner in liquid detergent and water and squeeze it dry in a clean cloth. Saturate it in engine oil. Wrap it in a clean absorbent cloth and squeeze to remove excess oil.
4. Clean cartridge by tapping gently on a flat surface. Replace if dirty.
5. Reinstall pre-cleaner, cartridge, and cover.

# Manual For John Deere 100 Series

**ML Yell**



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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical x 100 what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely becuase bromine is an ortho para director

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explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 °C

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture  $\text{HNO}_3$  3 M  $\text{H}_2\text{SO}_4$  in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by the nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion  $\text{NO}_2^+$  which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4.0 ml of nitric acid and 4.0 ml of sulfuric acid and will then react with 4.5 ml of bromobenzene this will produce o-bromonitrobenzene

and p bromonitrobenzene

**experiment 11 electrophilic aromatic substitution nitration** - Feb 10 2022

experimental spectra reaction a fluorobenzene stock fluorobenzene 1 h 1 h 19 f 19 f 13 c nmr fid for reference only sample 1 h 1 h 19 f 19 f 13 c nmr eas nitration of fluorobenzene crude product not available for submission for credit

**nitration of bromobenzene chem 154 gwu studocu** - Mar 26 2023

lab writeup the purpose of this lab was to observe where substituents are added on benzene ring way of electrophilic aromatic substitution the electrophilic

**5 nitration lab docx lab 5 nitration of bromobenzene** - Aug 19 2022

chem 2261 lab 5 nitration of bromobenzene introduction an electrophilic aromatic substitution reaction is an important process that allows for intermediate chemical structures to be used as precursors to pharmaceuticals and other industrially produced chemicals by adding different functional groups into an aromatic ring

lab report nitration of bromomobenzene ipl org - Feb 22 2023

lab report 10 nitration of bromobenzene raekwon filmore cm 244 section 40 march 27 2018 introduction for this experiment nitration of bromobenzene was the focus of the lab the benzene is an aromatic compound and when it reacts with wither a mixture of sulfuric acid or nitric acid creates what is known as a nitro group