

Answer all questions on this page

Questions 1-14: For each statement or question, record on your separate answer sheet the number of the correct response (one of three given, two completing the statement or answering the question). These questions are complete the rest of the 1981-82 Illinois Chemistry Examinations for High School Students.

1 Compared to the change of a positive ion, change of an electron is:

- (1) a positive magnitude and the same sign  
(2) a positive magnitude and the opposite sign  
(3) the same magnitude and the same sign  
(4) the same magnitude and the opposite sign

2 Which atom has the largest atomic radius?

- (1) potassium (K) (2) aluminum (Al)  
(3) sodium (Na) (4) cesium (Cs)

3 In the electrochemical series of the series, an element is defined as:

- (1) a metal which most probably tends to be oxidized  
(2) a metal of the most probable electron location  
(3) a metal which is oxidized by a positive metal ion  
(4) a metal which is oxidized by an electron around the nucleus

4 Which are correct statements for an atom, as it is the ground state, the electron:

- (1) relative energy as it moves to a higher energy state  
(2) relative energy as it moves to a lower energy state  
(3) relative energy as it moves to a higher energy state  
(4) relative energy as it moves to a lower energy state

5 Which polyatomic ion is found in the compound represented for the formula  $\text{NaHCO}_3$ ?

- (1) acetate (2) hydrogen sulfate  
(3) hydrogen carbonate (4) nitrate

6 The atomic mass of magnesium is the weighted average of the atomic masses of:

- (1) all of the naturally produced isotopes of Mg  
(2) all of the naturally occurring isotopes of Mg  
(3) the two most abundant naturally produced isotopes of Mg  
(4) the two most abundant naturally occurring isotopes of Mg

7 Which element has more than one known isotope?

- (1) hydrogen (2) neon  
(3) carbon (4) cesium

8 Two forms of water, ice and steam, are both made of the same chemical composition, but in the difference is their:

- (1) atomic structure  
(2) physical structure  
(3) relative abundance  
(4) percent composition

9 Which quantity can be calculated for a solid compound, given only the formula of the compound and the Periodic Table of the Elements?

- (1) the density of the compound  
(2) the heat of fusion of the compound  
(3) the melting point of each element in the compound  
(4) the percent composition by mass of each element in the compound

10 Which group closely agrees of chemical reactions?

- (1) decomposition and combination  
(2) decomposition and synthesis  
(3) decomposition and substitution  
(4) decomposition and synthesis

# June 2014 Living Environment Regents Answer Key

**LP Steffe**



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### **Table of Contents June 2014 Living Environment Regents Answer Key**

1. Understanding the eBook June 2014 Living Environment Regents Answer Key
  - The Rise of Digital Reading June 2014 Living Environment Regents Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying June 2014 Living Environment Regents Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook June 2014 Living Environment Regents Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from June 2014 Living Environment Regents Answer Key
  - Personalized Recommendations
  - June 2014 Living Environment Regents Answer Key User Reviews and Ratings
  - June 2014 Living Environment Regents Answer Key and Bestseller Lists

5. Accessing June 2014 Living Environment Regents Answer Key Free and Paid eBooks
  - June 2014 Living Environment Regents Answer Key Public Domain eBooks
  - June 2014 Living Environment Regents Answer Key eBook Subscription Services
  - June 2014 Living Environment Regents Answer Key Budget-Friendly Options
6. Navigating June 2014 Living Environment Regents Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - June 2014 Living Environment Regents Answer Key Compatibility with Devices
  - June 2014 Living Environment Regents Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of June 2014 Living Environment Regents Answer Key
  - Highlighting and Note-Taking June 2014 Living Environment Regents Answer Key
  - Interactive Elements June 2014 Living Environment Regents Answer Key
8. Staying Engaged with June 2014 Living Environment Regents Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers June 2014 Living Environment Regents Answer Key
9. Balancing eBooks and Physical Books June 2014 Living Environment Regents Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection June 2014 Living Environment Regents Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine June 2014 Living Environment Regents Answer Key
  - Setting Reading Goals June 2014 Living Environment Regents Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of June 2014 Living Environment Regents Answer Key
  - Fact-Checking eBook Content of June 2014 Living Environment Regents Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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