

GROUP 2 ELEMENTS AND THEIR COMPOUNDS

- show their outermost electron placed on a subshell
- readily react to lose 2 electrons
 - $\text{M} \rightarrow \text{M}^{2+} + 2\text{e}^-$

Trends down the group:

- 1st ionisation energy decreases
 - Distance between outermost electron and nucleus increases
 - More shielding
 - Counteractive effect of increasing nuclear charge
- Atomic radius increases
- Melting point generally decreases
 - Much lower than all transition metals

Physical properties:

- Good conductors of heat and electricity
 - Sea of delocalised electrons
- Form white compounds
 - Transition metals form coloured compounds
- Oxidation number of +2
- All compounds are toxic or almost toxic
 - Be compounds show some covalent character because small highly charged cation polarises anion shell of anion
- Oxides / hydroxides are basic
 - React with acids (proton acceptors)
 - Example:
 - BaO
 - $\text{BaO} + 2\text{H}^+ \rightarrow \text{Ba}^{2+} + \text{H}_2\text{O}$ which is exothermic
- Density = 2.83 g/cm^3
 - Average density of transition metals is about 7 g/cm^3
- Compared to group 2 metals:
 - Harder + denser
 - Higher melting point
 - Metallic bond is stronger (higher charge on cation & more electrons delocalised)

Chemical properties:

- Strong reducing agents
- Oxidation is losing electrons
- Reduction is gaining electrons:
 - Reducing agent - donates electrons (reducing other substances)
 - Oxidising agent - accepts electrons (oxidising itself)
- E.g.
 - $\text{ZnO} + \text{CO} \rightarrow \text{Zn} + \text{CO}_2$
 - $\text{ZnO} \rightarrow \text{Zn}^{2+} + 2\text{O}^{2-}$
 - $\text{CO} + 2\text{O}^{2-} \rightarrow \text{CO}_2^{2-}$

June 2013 Ocr Chemistry F321

JE Gale



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Table of Contents June 2013 Ocr Chemistry F321

1. Understanding the eBook June 2013 Ocr Chemistry F321
 - The Rise of Digital Reading June 2013 Ocr Chemistry F321
 - Advantages of eBooks Over Traditional Books
2. Identifying June 2013 Ocr Chemistry F321
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an June 2013 Ocr Chemistry F321
 - User-Friendly Interface
4. Exploring eBook Recommendations from June 2013 Ocr Chemistry F321
 - Personalized Recommendations
 - June 2013 Ocr Chemistry F321 User Reviews and Ratings
 - June 2013 Ocr Chemistry F321 and Bestseller Lists
5. Accessing June 2013 Ocr Chemistry F321 Free and Paid eBooks
 - June 2013 Ocr Chemistry F321 Public Domain eBooks
 - June 2013 Ocr Chemistry F321 eBook Subscription Services
 - June 2013 Ocr Chemistry F321 Budget-Friendly Options
6. Navigating June 2013 Ocr Chemistry F321 eBook Formats

- ePub, PDF, MOBI, and More
- June 2013 Ocr Chemistry F321 Compatibility with Devices
- June 2013 Ocr Chemistry F321 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of June 2013 Ocr Chemistry F321
 - Highlighting and Note-Taking June 2013 Ocr Chemistry F321
 - Interactive Elements June 2013 Ocr Chemistry F321
- 8. Staying Engaged with June 2013 Ocr Chemistry F321
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers June 2013 Ocr Chemistry F321
- 9. Balancing eBooks and Physical Books June 2013 Ocr Chemistry F321
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection June 2013 Ocr Chemistry F321
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine June 2013 Ocr Chemistry F321
 - Setting Reading Goals June 2013 Ocr Chemistry F321
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of June 2013 Ocr Chemistry F321
 - Fact-Checking eBook Content of June 2013 Ocr Chemistry F321
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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