Birla Institute of Technology and Science, Pilani (Rajasthan) First Semester 2022-23, December 21, 2022 Comprehensive Examination (Open Book) Part -II CHEM F214; Inorganic Chemistry I

Time: 2 hours 25 Mins

Max. Marks: 56

Q.1(a) Δ = (actual bond energy) - (energy for 100% covalent bond). Find out the Pauling electronegativities of C and Cl where the Δ values of C-H and H-Cl are 24.3 and 102.3 kJ/mol respectively [consider χ_H = 0]. If χ_H is shifted to 2, is there any advantage in the scale?

[3]

(b) Discuss the effect of aniline added to ammonia based on leveling effect.

(c) There exists a relationship between energy band gap and electronegativity of the elements present in a compound. What is the relation these hold? Can you provide some reasons for the existing relationship? [1+2]
(d) Calculate the equilibrium constant for the disproportionation reaction:

 $2Cu^{+}(aq) \iff Cu(s) + Cu^{2+}(aq) \text{ at } 298 \text{ K}$ $IE^{0}(Cu^{+}(Cu)) = 0.52 \text{ Y and } E^{0}(Cu^{2+}(Cu^{+})) = 0.15 \text{ Y} \cdot 2$

 $[E^0(Cu^*/Cu) = 0.52 \text{ V and } E^0(Cu^2*/Cu^*) = 0.15\text{V}; 2.303\text{RT/F} = 0.0591\text{V at } 298 \text{ K}].$

(e) Account for the fact that CH₂CH₂O is much more basic than CH₂COO.

[4] [2]

131

Q. 2 (a) How would you synthesize P(SiMe₃)₃ from Phosphorus? Draw the structure of P(SiMe₃)₃. Write down the expected product of the reaction between CH₃COCl and P(SiMe₃)₃.

[1+1+1]

(b) Account for the fact that in aqueous solution, lithium is as good a reducing agent as Cs despite of having higher ionization energy.
[3]

(c) With the mention of the products obtained when CCl₄ and SiCl₄ individually reacting with H₂O, provide justification to support the observations for these two reactions.
[4]

(d) What is the role of (n-Pr₄N)OH in the synthesis of ZSM-5? Which property of ZSM-5 will be tweaked by Si to Al ratio? If you want to introduce Fe²⁺ ions in ZSM-5, how would you proceed? [2+1+1]

Q. 3(a) To prepare I₃*, an oxidizing agent like bis(fluorosulpuryl)peroxide FSO₂OOSO₂F is used to react with I₂. Would you prefer H₂O₂ over FSO₂OOSO₂F for this reaction? Justify your answer. [3]

(b) Plenty of fluoride compounds of Xe are known. What could be the possible reason(s) for not having reports of formation of fluorides of Ne? If you imagine a Ne-fluoride compound, what would be the formula and what type of bonding is expected?
[2+2]

(c) Find out the symmetry point group of boron-phosphorus analogue of borazine B₂P₃((Ph)₆ [Consider the -Ph group as an atom] and mention the symmetry elements present. Amongst B₃P₃((Ph)₆ and borazine, who is more prone to non-planar structure and why?

[2+2]

(d) Comment on the aromaticity of S₄N₂ having a closed ring structure.

Q.4(a) Justify the participation of specific orbitals in π bonding from P and N atoms in trimeric phosphazene, $P_3N_3Cl_6$. With the typical sets of orbitals once chosen from P and N centers, discuss the 'island model' for the molecule $P_3N_3Cl_6$. [2+2]

(b) Draw the molecular orbitals of the 3c-2e bond considering a hypothetical model with the atoms as H-B-H in B₂H₆. Why is the 3c-2e bonds on B-H-B system justified over 3c-2e bond on H-B-H system? [3+2]

(c) Explain the bond order of Re-Re bond in [Re₂Cl₈]⁺. Justify the arrangement of chlorine atoms around each Re atom relatively and the geometry of each Re atom. [2+3]

******END*****

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Table of Contents June 2013 F214 Past Paper

- 1. Understanding the eBook June 2013 F214 Past Paper
 - The Rise of Digital Reading June 2013 F214 Past Paper
 - Advantages of eBooks Over Traditional Books
- 2. Identifying June 2013 F214 Past Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an June 2013 F214 Past Paper
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from June 2013 F214 Past Paper
 - Personalized Recommendations
 - June 2013 F214 Past Paper User Reviews and Ratings
 - ∘ June 2013 F214 Past Paper and Bestseller Lists

- 5. Accessing June 2013 F214 Past Paper Free and Paid eBooks
 - o June 2013 F214 Past Paper Public Domain eBooks
 - o June 2013 F214 Past Paper eBook Subscription Services
 - June 2013 F214 Past Paper Budget-Friendly Options
- 6. Navigating June 2013 F214 Past Paper eBook Formats
 - o ePub, PDF, MOBI, and More
 - June 2013 F214 Past Paper Compatibility with Devices
 - June 2013 F214 Past Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of June 2013 F214 Past Paper
 - Highlighting and Note-Taking June 2013 F214 Past Paper
 - Interactive Elements June 2013 F214 Past Paper
- 8. Staying Engaged with June 2013 F214 Past Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers June 2013 F214 Past Paper
- 9. Balancing eBooks and Physical Books June 2013 F214 Past Paper
 - Benefits of a Digital Library
 - ∘ Creating a Diverse Reading Collection June 2013 F214 Past Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine June 2013 F214 Past Paper
 - Setting Reading Goals June 2013 F214 Past Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of June 2013 F214 Past Paper
 - Fact-Checking eBook Content of June 2013 F214 Past Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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