

1.	(a)	(i)	electrons ...16.... protons16 neutrons ...16.... All correct (1)	1	[8]
		(ii)	they are isotopes (1) The word 'isotope' or some derivative is essential	1	
		(iii)	S^{2-} / sulphide(2-) ion / sulphur -2 ion / sulphur with two extra electrons (1) but not just S or sulphur	1	
	(b)	(i)	$X(g) + e^- \rightarrow X^-(g)$ ($S(g) + e^- \rightarrow S^-(g)$) (1) $X(g) + e^- \rightarrow X^{2-}(g)$ ($S(g) + e^- \rightarrow S^{2-}(g)$) (1) can have Z in place of X ²⁻ but not Z ²⁻ Penalize omission of state symbol (g) only once.	2	
		(ii)	adding a negative onto a <u>negative species</u> requires energy to overcome repulsion (1) or second electron repelled by negative charge Not second electron repelled by first	1	
		(c)	(Bdecay from) Y produces only one isotope(^{39}Cl) (1) natural sample contains two isotopes (35 and 37) (1)	2	
	2.	(i)	magnesium / lead / iron (1)	1	
		(ii)	iron (1)	1	
		(iii)	carbon / phosphorus (1)	1	
		(iv)	magnesium / lead (1)	1	
		(v)	carbon / lead (1)	1	
		(b)	Helium (1) because it is the smallest atom (1) with an unshielded nucleus (1) Hydrogen can score max 1 for 'unshielded nucleus' If candidate chooses Carbon, i.e. uses list from part (a), can score 1 mark for reference to 'smallest atom'	3	
		(c)	(i)	ability of a positive ion / cation to attract electrons / distort / pull the electron cloud (1) around an anion (1)	2
			(ii)	size (1) Not atomic radius charge (1) charge density worth (2) volume of ion can score in place of size	2
			(iii)	magnesium chloride magnesium ion smaller than calcium ion (1) polarizes / distorts / attracts electrons around chloride more (1)	2
					[14]

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Staley Wilbraham (Matta, Waterman)



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