

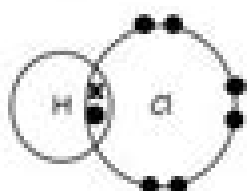
# C3 - STRUCTURE AND BONDING 3.5 Covalent bond

## Aiming for Grade 4

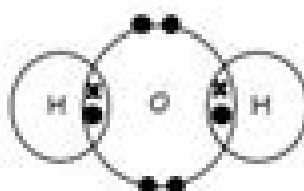
Ionic bonds form between **metals** and **non-metals**. Covalent bonds form between **non-metals**. Covalent bonds form when atoms **share** pairs of **electrons**. These bonds are very **strong**.

**Share, Metals, strong, non-metals, electrons**  
You can use the words more than once.

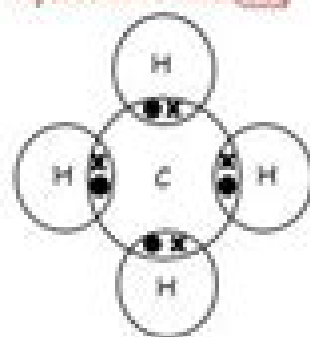
Name these small molecules and state their formulae.



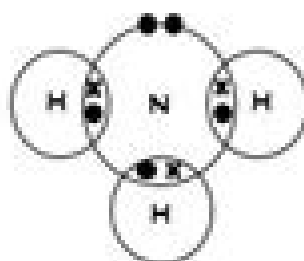
Hydrochloric Acid ( $\text{HCl}$ )



Water ( $\text{H}_2\text{O}$ )



Methane ( $\text{CH}_4$ )



Ammonia ( $\text{NH}_3$ )

Assessment

## Aiming for Grade 6

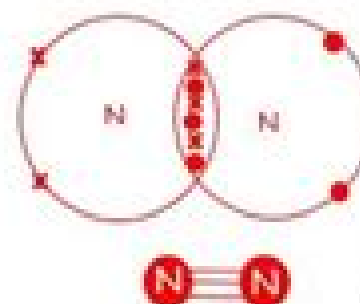
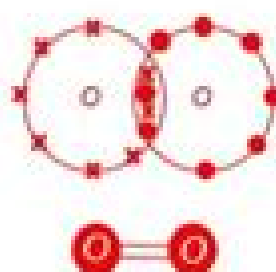
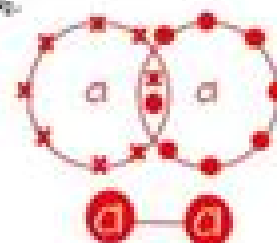
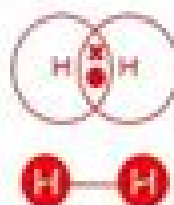
Describe how Hydrogen and Chlorine form a covalent bond to form Hydrochloric acid.

To gain full outer shells Hydrogen needs to gain an electron and Chlorine needs to lose one. Hydrogen shares its outer shell electron with Chlorine and Chlorine shares one of its electrons.

What is a double covalent bond?

Two shared pairs of electrons

Draw dot and cross diagrams and ball and stick diagrams for  $\text{H}_2$ ,  $\text{Cl}_2$ ,  $\text{O}_2$ , and  $\text{N}_2$ .



Assessment

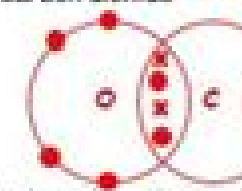
## Aiming

Draw ball and stick or the following covalent

Three hydrogen atoms  
one phosphorus (phos)



Carbon dioxide



Ethane (2 carbon with hydrogen)



How do double and triple properties?

A double or triple bond the bond making the me

Assessment  
Grade 6

1. I can draw dot and

# June 2013 Aqa Chemistry C3 Mark Scheme

**Richard Bailey**



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