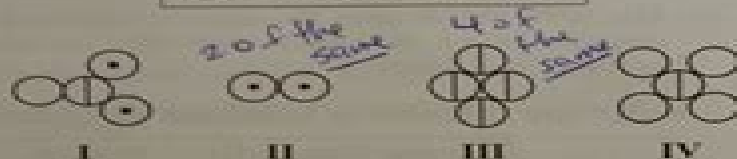
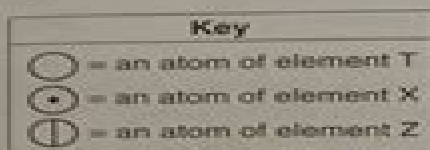


39 Given four particle models:



Which two models can be classified as elements?

- (1) I and II
 (2) I and IV
 (3) II and III
 (4) II and IV

40 After being thoroughly stirred at $10.^{\circ}\text{C}$, which mixture is heterogeneous?

- (1) 25.0 g of KCl and 100. g of H_2O *unsat.*
 (2) 25.0 g of KNO_3 and 100. g of H_2O *supersat.*
 (3) 25.0 g of NaCl and 100. g of H_2O *unsat.*
 (4) 25.0 g of NaNO_3 and 100. g of H_2O *unsat.*

41 Which two compounds are electrolytes?

- (1) KOH and CH_3COOH
 (2) KOH and C_3H_{12}
 (3) CH_3OH and CH_3COOH
 (4) CH_3OH and C_3H_{12}

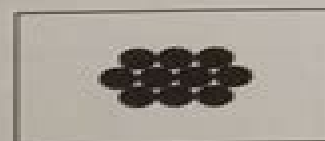
42 Which statement explains why a CO_2 molecule is nonpolar?

- (1) Carbon and oxygen are both nonmetals.
 (2) Carbon and oxygen have different electronegativities.
 (3) The molecule has a symmetrical distribution of charge.
 (4) The molecule has an asymmetrical distribution of charge.

43 Which temperature change indicates an increase in the average kinetic energy of the molecules in a sample?

- (1) $15.^{\circ}\text{C}$ to 298 K $\uparrow$
 (2) $37.^{\circ}\text{C}$ to 273 K $\downarrow$
 (3) 305 K to $0.^{\circ}\text{C}$ $\downarrow$
 (4) 355 K to $25.^{\circ}\text{C}$ $\downarrow$

44 Given the particle diagram:



Which substance at STP can be represented by this particle diagram? *Solid*

- (1) N_2 *(g)*
 (2) H_2 *(g)*
 (3) Mg *(s)*
 (4) Kr *(g)*

45 Which type of equilibrium exists in a sealed flask containing $\text{Br}_2(\ell)$ and $\text{Br}_2(\text{g})$ at 298 K and 1.0 atm ?

- (1) static phase equilibrium
 (2) static solution equilibrium
 (3) dynamic phase equilibrium
 (4) dynamic solution equilibrium

46 What are the products when potassium hydroxide reacts with hydrochloric acid?

- (1) $\text{KH}(\text{s})$, $\text{Cl}^+(\text{aq})$, and $\text{OH}^-(\text{aq})$
 (2) $\text{K}(\text{s})$, $\text{Cl}_2(\text{g})$, and $\text{H}_2\text{O}(\ell)$
 (3) $\text{KCl}(\text{aq})$ and $\text{H}_2\text{O}(\ell)$
 (4) $\text{KOH}(\text{aq})$ and $\text{Cl}_2(\text{g})$

47 In a titration, 20.0 milliliters of a 0.150 M $\text{NaOH}(\text{aq})$ solution exactly neutralizes 24.0 milliliters of an $\text{HCl}(\text{aq})$ solution. What is the concentration of the $\text{HCl}(\text{aq})$ solution?

- (1) 0.125 M
 (2) 0.180 M
 (3) 0.250 M
 (4) 0.360 M

48 What fraction of a Sr-90 sample remains unchanged after 87.3 years?

- (1) $\frac{1}{2}$
 (2) $\frac{1}{3}$
 (3) $\frac{1}{4}$
 (4) $\frac{1}{8}$

Handwritten calculation:

$$3 = \frac{87.3}{29.14}$$

$$\frac{1}{2} = \frac{1}{8}$$

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