

Part B-1

Answer all questions in this part.

Directions (31–50): For each statement or question, record on your separate answer sheet the number of the word or expression that, of those given, best completes the statement or answers the question. Some questions may require the use of the 2011 Edition Reference Tables for Physical Setting/Chemistry.

- 31 What is the number of electrons in an Al^{3+} ion? Low 5
 (1) 10 (3) 3
 (2) 13 (4) 16 Al +13p

- 32 The valence electron of which atom in the ground state has the greatest amount of energy?
 (1) cesium 23.5 (3) rubidium
 (2) lithium 23.3 (4) sodium

- 33 The numbers of protons and neutrons in each of four different atoms are shown in the table below.

Protons and Neutrons in Four Different Atoms

Atom	Number of Protons	Number of Neutrons
A	8	8
D	9	9
E	9	10
G	10	10

Which two atoms represent isotopes of the same element?

- (1) A and D (3) E and D
 (2) A and G (4) E and G

- 34 Which elements have the most similar chemical properties?

- (1) boron and carbon
 (2) oxygen and sulfur
 (3) aluminum and bromine
 (4) argon and silicon

- 35 Which element reacts with oxygen to form ionic bonds?

- (1) calcium (3) chlorine
 (2) hydrogen (4) nitrogen

- 36 The table below gives the atomic mass and the abundance of the two naturally occurring isotopes of chlorine.

Naturally Occurring Isotopes of Chlorine

Isotopes	Atomic Mass of the isotope (u)	Natural Abundance (%)
^{35}Cl	34.97	75.76
^{37}Cl	36.97	24.24

Which numerical setup can be used to calculate the atomic mass of the element chlorine?

- (1) $(34.97 \text{ u})(75.76) + (36.97 \text{ u})(24.24)$
 (2) $(34.97 \text{ u})(0.2424) + (36.97 \text{ u})(0.7576)$
 (3) $(34.97 \text{ u})(0.7576) + (36.97 \text{ u})(0.2424)$
 (4) $(34.97 \text{ u})(24.24) + (36.97 \text{ u})(75.76)$

- 37 Which general trends in first ionization energy and electronegativity values are demonstrated by Group 15 elements as they are considered in order from top to bottom?

- (1) The first ionization energy decreases and the electronegativity decreases.
 (2) The first ionization energy increases and the electronegativity increases.
 (3) The first ionization energy decreases and the electronegativity increases.
 (4) The first ionization energy increases and the electronegativity decreases.

- 38 An aluminum sample has a mass of 80.01 g and a density of 2.70 g/cm³. According to the data, to what number of significant figures should the calculated volume of the aluminum sample be expressed?

- (1) 1 (3) 3
 (2) 2 (4) 4

$$D = \frac{M}{V} \quad ; \quad V = \frac{M}{D}$$

$$\frac{80.01 \text{ (4)}}{2.70 \text{ (3)}} \text{ (3)}$$

June 2015 Chemistry Regents Answers

Julia Schneider



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Table of Contents June 2015 Chemistry Regents Answers

1. Understanding the eBook June 2015 Chemistry Regents Answers
 - The Rise of Digital Reading June 2015 Chemistry Regents Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying June 2015 Chemistry Regents Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook June 2015 Chemistry Regents Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from June 2015 Chemistry Regents Answers
 - Personalized Recommendations
 - June 2015 Chemistry Regents Answers User Reviews and Ratings
 - June 2015 Chemistry Regents Answers and Bestseller Lists
5. Accessing June 2015 Chemistry Regents Answers Free and Paid eBooks
 - June 2015 Chemistry Regents Answers Public Domain eBooks
 - June 2015 Chemistry Regents Answers eBook Subscription Services
 - June 2015 Chemistry Regents Answers Budget-Friendly Options
6. Navigating June 2015 Chemistry Regents Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - June 2015 Chemistry Regents Answers Compatibility with Devices
 - June 2015 Chemistry Regents Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of June 2015 Chemistry Regents Answers
 - Highlighting and Note-Taking June 2015 Chemistry Regents Answers
 - Interactive Elements June 2015 Chemistry Regents Answers
8. Staying Engaged with June 2015 Chemistry Regents Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers June 2015 Chemistry Regents Answers
- 9. Balancing eBooks and Physical Books June 2015 Chemistry Regents Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection June 2015 Chemistry Regents Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine June 2015 Chemistry Regents Answers
 - Setting Reading Goals June 2015 Chemistry Regents Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of June 2015 Chemistry Regents Answers
 - Fact-Checking eBook Content of June 2015 Chemistry Regents Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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