

LONDON COLLEGE OF INTERNATIONAL CHEMISTS – 6000 (004)

NAME: _____ COURSE: _____

ENTRY NUMBER: _____ ALPH. NO. _____

IC: 0000 (CONCEPTS IN ORGANIC CHEMISTRY)

0000 00

Class : 0000000000
 Date : 0000 - 0000

Page: No. _____

Time : 000 Minutes

Answer ALL the questions.

PART – A

10 x 2 = 20 marks

1. State three bonding/intermolecular principles.
2. Describe hydrogen bonding in alcohols in the formula CH_3OH .
3. Write the chemical formula for $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$.
4. Describe the structure of hydrogen in H_2 gas.
5. What is the relationship between the number of valence electrons of an element and the number of bonds it can form?
6. Describe and explain how the number of valence electrons of an element affects its chemical properties.
7. What are ionic bonds? Give an example.
8. What are covalent bonds? Give an example.
9. Describe the structure of water (H_2O).
10. What are ionic compounds? Give an example.

Answer any THREE questions.

PART – B

3 x 5 = 15 marks

11. What are ionic and covalent bonds? Describe the difference between them.
12. Explain the difference between the chemical properties of ionic and covalent compounds.
13. Describe the structure of water (H_2O) and explain its properties.
14. Describe the structure of hydrogen (H_2) and explain its properties.
15. Describe the structure of methane (CH_4) and explain its properties.
16. Describe the structure of ammonia (NH_3) and explain its properties.
17. Describe the structure of hydrogen chloride (HCl) and explain its properties.
18. Describe the structure of hydrogen sulfide (H_2S) and explain its properties.
19. Describe the structure of hydrogen cyanide (HCN) and explain its properties.
20. Describe the structure of hydrogen cyanide (HCN) and explain its properties.
21. Describe the structure of hydrogen cyanide (HCN) and explain its properties.
22. Describe the structure of hydrogen cyanide (HCN) and explain its properties.

PART – C

2 x 5 = 10 marks

Answer any TWO questions.

23. What are ionic and covalent bonds? Describe the difference between them.
24. Explain the difference between the chemical properties of ionic and covalent compounds.
25. Describe the structure of water (H_2O) and explain its properties.
26. Describe the structure of hydrogen (H_2) and explain its properties.
27. Describe the structure of methane (CH_4) and explain its properties.
28. Describe the structure of ammonia (NH_3) and explain its properties.
29. Describe the structure of hydrogen chloride (HCl) and explain its properties.
30. Describe the structure of hydrogen sulfide (H_2S) and explain its properties.
31. Describe the structure of hydrogen cyanide (HCN) and explain its properties.
32. Describe the structure of hydrogen cyanide (HCN) and explain its properties.

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