

28. Which of the following compounds is square planar and does not have any unpaired electron?
 (a) $\text{Ni}(\text{CO})_4$ (b) $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$
 (c) $[\text{NiCl}_4]^{2-}$ (d) $[\text{Ni}(\text{CN})_4]^{2-}$
29. White anhydrous copper sulphate decomposes to give
 (a) $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ (b) $\text{CuSO}_4 \cdot \text{H}_2\text{O}$
 (c) $\text{CuO} + \text{SO}_2$ (d) Cu
30. If NaOH is added to an aqueous solution of zinc ions, a white precipitate appears and on adding excess of NaOH , the precipitate dissolves. In this solution zinc exist in the
 (a) cationic part
 (b) anionic part
 (c) both in cationic and anionic part
 (d) there is no zinc in the solution
31. Chlorine gas is dried over
 (a) CaO (b) NaOH
 (c) KOH (d) conc. H_2SO_4
32. Which of the following halogens does not form its oxyacids?
 (a) Fluorine (b) Chlorine
 (c) Bromine (d) Iodine
33. Which of the following hydrides has the lowest boiling point?
 (a) H_2O (b) H_2S
 (c) H_2Se (d) H_2Te
34. Glass is a
 (a) polymeric mixture
 (b) microcrystalline solid
 (c) supercooled liquid
 (d) gel
35. Thermite is a mixture of iron oxide and
 (a) zinc powder (b) sodium shavings
 (c) potassium metal (d) aluminium powder
36. The volume of water to be added to 100 cm^3 of $0.5 \text{ N H}_2\text{SO}_4$ to get decinormal concentration is
 (a) 400 cm^3 (b) 500 cm^3
 (c) 450 cm^3 (d) 100 cm^3
37. The rate of a first order reaction is $1.5 \times 10^{-2} \text{ mol L}^{-1} \text{ min}^{-1}$ at 0.5 M concentration of the reactant. The half-life of the reaction is
 (a) 7.53 min (b) 0.383 min
 (c) 23.1 min (d) 8.73 min
38. The hydrolysis of ester in alkaline medium is a
 (a) first order reaction with molecularity 1
 (b) second order reaction with molecularity 2
 (c) first order reaction with molecularity 2
 (d) second order reaction with molecularity 1
39. Which is correct statement?
 (a) Starch is a polymer of α -glucose
 (b) Amylose is a component of cellulose
 (c) Proteins are compounds of only one type of amino acid
 (d) In cyclic structure of fructose, there are four carbons and one oxygen atom
40. Oleic, stearic and palmitic acids are
 (a) nucleic acids (b) amino acids
 (c) fatty acids (d) None of these
41. Arrange the following halides in the decreasing order of $\text{S}_{\text{N}}1$ reactivity.
 $\text{CH}_3\text{CH}_2\text{CH}_2\text{Cl}$ (I) $\text{CH}_2=\text{CHCH}(\text{Cl})\text{CH}_3$ (II)
 $\text{CH}_3\text{CH}_2\text{CH}(\text{Cl})\text{CH}_3$ (III)
 (a) $\text{I} > \text{II} > \text{III}$ (b) $\text{II} > \text{I} > \text{III}$
 (c) $\text{II} > \text{III} > \text{I}$ (d) $\text{III} > \text{II} > \text{I}$
42. Catalyst used in the dimerisation of acetylene to prepare chloroprene is
 (a) $\text{HgSO}_4 + \text{H}_2\text{SO}_4$ (b) Cu_2Cl_2
 (c) $\text{Cu}_2\text{Cl}_2 + \text{NH}_4\text{Cl}$ (d) $\text{Cu}_2\text{Cl}_2 + \text{NH}_4\text{OH}$
43. The phenomenon of radioactivity is associated with
 (a) decay of nucleus
 (b) fusion of nucleus
 (c) emission of electrons or protons
 (d) rearrangement in the extra nuclear electron
44. If Z is the number of atoms in the unit cell that represents the closest packing sequence - ABC, ABC -, the number of tetrahedral voids in the unit cell is equal to
 (a) Z (b) $2Z$
 (c) $\frac{Z}{2}$ (d) $\frac{Z}{4}$
45. Conjugate acid of SO_4^{2-} is
 (a) HSO_4^- (b) HSO_4
 (c) H_2SO_4 (d) SO_4^-
46. $\text{AgCl} + \text{KI} \rightleftharpoons \text{KCl} + \text{AgI}$
 In the above reaction as KI is added, the equilibrium is shifted towards right giving more AgI precipitate, because

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