

Week 10: Precipitation Reactions

1. Start *Virtual ChemLab*, select *Reactions and Stoichiometry*, and then select *Precipitation Reactions* from the list of assignments. The lab will open in the Inorganic laboratory.
2. React each of the cations (across the top) with each of the anions (down the left) according to the data table using the following procedures:

Data Table

	AgNO_3 (Ag^+)	$\text{Pb}(\text{NO}_3)_2$ (Pb^{2+})	$\text{Ca}(\text{NO}_3)_2$ (Ca^{2+})
Na_2CO_3 (CO_3^{2-})	a cloudy peach	f milky white color	k milky white color
Na_2S (S^{2-})	b cloudy black color	g cloudy black color	l no visible reaction
NaOH (OH^-)	c cloudy brownish	h no visible reaction	m milky white color
Na_2SO_4 (SO_4^{2-})	d no visible reaction	i milky white color	n no visible reaction
NaCl (Cl^-)	e pale white color	j milky white color	o no visible reaction

- a. Drag a test tube from the box and place it on the metal test tube stand. You can then click on the bottle of Ag^+ ion solution on the shelf to add it to the test tube.
- b. Click on the *Divide* button on the edge of the lab bench four times to make four additional test tubes containing Ag^+ . With one test tube in the metal stand and four others in the blue rack, click on the Na_2CO_3 bottle on the reagent shelf and observe what happens in the window at the top left. Record your observation in the table above. If the solution remains clear, record NR for no reaction. Drag this test tube to the red disposal bucket on the right.
- c. Drag a second tube from the blue rack to the metal stand. Add Na_2S , record your observations and discard the tube. Continue with the third, fourth and fifth tube, but add NaOH , Na_2SO_4 , and NaCl respectively. Record your observations in the table above.

Investigating Precipitation Reactions Lab Answers

Sam Stuart



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