Section 18.4 Neutralization

In your textbook, read about neutralization and titration.

For each item in Column A, write the letter of the matching item in Column B.

Column A.

- A chemical dye that changes color based on the pH of a solution
 - A method for using a neutralization reaction to determine the concentration of a solution
- A reaction in which an acid and a base react to produce a salt and water
- 4. A solution of known concentration
 - 5. An ionic product of an acid-base reaction
- The point in a titration in which an indicator changes color
- 7. The stoichiometric point of a titration

Column B

- a. acid-base indicator
- b. end point
- c. equivalence point
- d. neutralization
- e. sait.
- f. standard solution
- g. titration

Complete the following table, indicating the formula and name of the salt formed by a neutralization reaction between the listed acid and base.

Acid	Base	Salt formula	Saft name
8. HCI	кон	KCI	potassium chloride
9. H ₂ SO ₄	Mg(0H) ₂		
10. H ₃ PO ₄	NaOH		
11. HNO ₃	Fe(OH) ₃		
12. H ₃ PO ₄	Ca(OH0 ₂		

In the space at the left, write I through 4 to show the correct sequence of the steps in performing a titration using a pH meter. Then, write 5 through 8 to sequence the steps used to calculate the concentration of the unknown solution.

Sequence of Steps

- Continue adding the standard solution to the solution of unknown concentration until the equivalence point is reached.
 - 14. Fill a buret with the standard solution.
- 15. Start adding the standard solution slowly, with mixing, to the solution of unknown concentration, reading the pH at regular intervals.

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