

**CHAPTER 6 REVIEW***Chemical Bonding***SECTION 4****SHORT ANSWER** Answer the following questions in the space provided.

1.   b   In metals, the valence electrons are considered to be  
(a) attached to particular positive ions. (c) immobile.  
(b) shared by all surrounding atoms. (d) involved in covalent bonds.
2.   a   The fact that metals are malleable and ionic crystals are brittle is best explained in terms of their  
(a) chemical bonds. (c) enthalpies of vaporization.  
(b) London forces. (d) polarity.
3.   d   As light strikes the surface of a metal, the electrons in the electron sea  
(a) allow the light to pass through.  
(b) become attached to particular positive ions.  
(c) fall to lower energy levels.  
(d) absorb and re-emit the light.
4.   d   Mobile electrons in the metallic bond are responsible for  
(a) luster. (c) electrical conductivity.  
(b) thermal conductivity. (d) All of the above.
5.   a   In general, the strength of the metallic bond \_\_\_\_\_ moving from left to right on any row of the periodic table.  
(a) increases (c) remains the same  
(b) decreases (d) varies
6.   c   When a metal is drawn into a wire, the metallic bonds  
(a) break easily. (c) do not break.  
(b) break with difficulty. (d) become ionic bonds.
7. Use the concept of electron configurations to explain why the number of valence electrons in metals tends to be less than the number in most nonmetals.

Most metals have their outer electrons in s orbitals, while nonmetals have their  
outer electrons in p orbitals.

# Introduction To Chemical Bonding Review Answers

**Brendan G. Carr**



## **Introduction To Chemical Bonding Review Answers:**

## Adopting the Song of Expression: An Mental Symphony within **Introduction To Chemical Bonding Review Answers**

In a world eaten by screens and the ceaseless chatter of instant transmission, the melodic elegance and psychological symphony created by the written term usually diminish in to the background, eclipsed by the persistent noise and disturbances that permeate our lives. However, located within the pages of **Introduction To Chemical Bonding Review Answers** a charming fictional prize brimming with fresh feelings, lies an immersive symphony waiting to be embraced. Constructed by a wonderful composer of language, this charming masterpiece conducts viewers on a mental journey, skillfully unraveling the hidden melodies and profound affect resonating within each carefully crafted phrase. Within the depths with this emotional examination, we can discover the book is key harmonies, analyze their enthralling writing fashion, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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