

Three organic compounds contain carbon, hydrogen and oxygen only.

(a) An organic compound B has the following composition by mass:

$$\text{C, 66.7\%}; \text{H, 11.1\%}; \text{O, 22.2\%}$$

Calculate the empirical formula of compound B.

$$\text{C} \rightarrow \frac{66.7}{12} = 5.558$$

$$\text{H} \rightarrow \frac{11.1}{1} = 11.1$$

$$\text{O} \rightarrow \frac{22.2}{16} = 1.387$$

$$\text{C} \rightarrow \frac{5.558}{1.387} = 4$$

$$\text{H} \rightarrow \frac{11.1}{1.387} = 8$$

$$\text{O} \rightarrow \frac{1.387}{1.387} = 1$$

$$\text{empirical formula} = \text{C}_4\text{H}_8\text{O}$$

(b) Compound B has the empirical formula $\text{C}_4\text{H}_8\text{O}$ and a relative molecular mass of 88.

Calculate the molecular formula of compound B.

$$\text{M}_{\text{empirical}} = 72$$

$$\text{M}_{\text{molecular}} = 88$$

$$n = \frac{88}{72} = 1.22$$

$$\frac{\text{relative mass}}{\text{empirical mass}} = \frac{88}{72} = 1.22 \quad \text{molecular formula} = \text{C}_5\text{H}_{10}\text{O}_2$$

$$\left[\begin{array}{l} 1 \times \text{C}_4\text{H}_8\text{O} = \text{C}_4\text{H}_8\text{O} \\ \text{empirical formula} \end{array} \right]$$

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