

# Key

## Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, " $PerV = nRT$ ", and the universal gas constant  $R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$  to solve the following problems:

If pressure is needed in kPa then convert by multiplying by  $101.3 kPa / 1 atm$  to get  
 $R = 8.31 kPa \cdot L / (K \cdot mole)$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 atm)(12 L)}{4 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol}}$$

$$T = 204.63 K$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of  $87^{\circ}C$ , how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 atm \cdot 31 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 360 K}$$

$$n = 1.2586 mol$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 400 K}{60 L}$$

$$P = 1.642 atm$$

$$or \quad P = 166.29 kPa$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of  $56^{\circ}C$ , what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 327 K}{0.09 atm}$$

$$V = 2310.93 L$$

- 5) If I have 17 moles of gas at a temperature of  $67^{\circ}C$ , and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 390 K}{88.89 L}$$

$$P = 5.34 atm$$

$$or \quad P = 540.61 kPa$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 atm \cdot 25 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 300 K}$$

$$n = 0.5075 mol$$

# Ideal Gas Law Answers Pg 24

**Team Prabhat**



## **Ideal Gas Law Answers Pg 24:**

**Physics** John D. Cutnell, Kenneth W. Johnson, 2009-09-08 Designed for medical professionals who may struggle with making the leap to conceptual understanding and applying physics the eighth edition continues to build transferable problem solving skills It includes a set of features such as Analyzing Multiple Concept Problems Check Your Understanding Concepts Calculations and Concepts at a Glance This helps the reader to first identify the physics concepts then associate the appropriate mathematical equations and finally to work out an algebraic solution Oswaal NCERT Exemplar (Problems - Solutions) Class 11 Physics, Chemistry and Biology (Set of 3 Books) For 2024 Exam Oswaal Editorial Board, 2023-10-28 Description of the product Chapter wise and Topic wise presentation Chapter wise Objectives A sneak peek into the chapter Mind Map A single page snapshot of the entire chapter Revision Notes Concept based study materials Tips Tricks Useful guidelines for attempting each question perfectly Some Commonly Made Errors Most common and unidentified errors are focused Expert Advice Oswaal Expert Advice on how to score more Oswaal QR Codes For Quick Revision on your Mobile Phones and Tablets Oswaal NCERT Exemplar (Problems - Solutions) Class 11 Physics, Chemistry and Mathematics (Set of 3 Books) For 2024 Exam Oswaal Editorial Board, 2023-10-28 Description of the product Chapter wise and Topic wise presentation Chapter wise Objectives A sneak peek into the chapter Mind Map A single page snapshot of the entire chapter Revision Notes Concept based study materials Tips Tricks Useful guidelines for attempting each question perfectly Some Commonly Made Errors Most common and unidentified errors are focused Expert Advice Oswaal Expert Advice on how to score more Oswaal QR Codes For Quick Revision on your Mobile Phones and Tablets **Chemistry** Richard Post, Chad Snyder, Clifford C. Houk, 2020-08-20 A practical complete and easy to use guide for understanding major chemistry concepts and terms Master the fundamentals of chemistry with this fast and easy guide Chemistry is a fundamental science that touches all other sciences including biology physics electronics environmental studies astronomy and more Thousands of students have successfully used the previous editions of Chemistry Concepts and Problems A Self Teaching Guide to learn chemistry either independently as a refresher or in parallel with a college chemistry course This newly revised edition includes updates and additions to improve your success in learning chemistry This book uses an interactive self teaching method including frequent questions and study problems increasing both the speed of learning and retention Monitor your progress with self tests and master chemistry quickly This revised Third Edition provides a fresh step by step approach to learning that requires no prerequisites lets you work at your own pace and reinforces what you learn ensuring lifelong mastery Master the science of basic chemistry with this innovative self paced study guide Teach yourself chemistry refresh your knowledge in preparation for medical studies or other coursework or enhance your college chemistry course Use self study features including review questions and quizzes to ensure that you re really learning the material Prepare for a career in the sciences medicine or engineering with the core content in this user friendly guide Authored by expert postsecondary

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