

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, "PerV-nRT", and the universal gas constant $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by $101.3 \text{ kPa} / 1 \text{ atm}$ to get
 $R = 8.31 \text{ kPa} \cdot \text{L} / (\text{K} \cdot \text{mole})$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 \text{ atm})(12 \text{ L})}{4 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}}$$

$$T = 204.63 \text{ K}$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87°C , how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 \text{ atm} \cdot 31 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 360 \text{ K}}$$

$$n = 1.2586 \text{ mol}$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 400 \text{ K}}{60 \text{ L}}$$

$$P = 1.642 \text{ atm}$$

$$\text{or } P = 166.29 \text{ kPa}$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56°C , what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 327 \text{ K}}{0.09 \text{ atm}}$$

$$V = 2310.93 \text{ L}$$

- 5) If I have 17 moles of gas at a temperature of 67°C , and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 390 \text{ K}}{88.89 \text{ L}}$$

$$P = 5.34 \text{ atm}$$

$$\text{or } P = 540.61 \text{ kPa}$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 \text{ atm} \cdot 25 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 300 \text{ K}}$$

$$n = 0.5075 \text{ mol}$$

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Holt Chemistry File, 1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket

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one two and three dimensions Included in the two dimensional program are Maxwell's equations and thermal and radiation diffusion The programs were developed by the author during the years 1952-1985 at the Lawrence Livermore National Laboratory The largest main frame computers available in the early 1950s were required to solve hydrodynamic problems in one space dimension by using forty mass points Subsequently numerical methods were developed for solving problems in two and three space dimensions but application of these methods had to wait until the main frame computers were large enough to tackle meaningful problems At the present time lap top computers can use these methods to solve problems in three space dimensions with the detail of 10 000 mass points The numerical procedures described in the text permit the exact conservation of physical properties in the solutions of the fundamental laws of mechanics 1 conservation of mass 2 conservation of momentum 3 conservation of energy The laws of mechanics are universal in their application Examples are given for the same computer simulation programs solving problems of penetration mechanics surface waves from earthquakes shock waves in solids and gases failure of materials **Oil & Gas Journal**, 1923 Practical Engineer, 1896

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 - The Rise of Digital Reading Ideal Gas Additional Problems Holt
 - Advantages of eBooks Over Traditional Books
2. Identifying Ideal Gas Additional Problems Holt
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ideal Gas Additional Problems Holt
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ideal Gas Additional Problems Holt
 - Personalized Recommendations
 - Ideal Gas Additional Problems Holt User Reviews and Ratings
 - Ideal Gas Additional Problems Holt and Bestseller Lists
5. Accessing Ideal Gas Additional Problems Holt Free and Paid eBooks
 - Ideal Gas Additional Problems Holt Public Domain eBooks

- Ideal Gas Additional Problems Holt eBook Subscription Services
- Ideal Gas Additional Problems Holt Budget-Friendly Options
- 6. Navigating Ideal Gas Additional Problems Holt eBook Formats
 - ePub, PDF, MOBI, and More
 - Ideal Gas Additional Problems Holt Compatibility with Devices
 - Ideal Gas Additional Problems Holt Enhanced eBook Features
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 - Interactive Elements Ideal Gas Additional Problems Holt
- 8. Staying Engaged with Ideal Gas Additional Problems Holt
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ideal Gas Additional Problems Holt
- 9. Balancing eBooks and Physical Books Ideal Gas Additional Problems Holt
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ideal Gas Additional Problems Holt
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ideal Gas Additional Problems Holt
 - Setting Reading Goals Ideal Gas Additional Problems Holt
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ideal Gas Additional Problems Holt
 - Fact-Checking eBook Content of Ideal Gas Additional Problems Holt
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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