

Test Review No 2

Early Theories of Matter.

Democritus (Greek Philosopher ~460 BC) proposed all that matter is composed of particles called atoms (Greek for "uncuttable"). He envisioned atoms of different substances as having different geometric shapes. The idea did not gain acceptance.

Aristotle proposed that there were four elements: earth, air, fire, and water. This idea gained acceptance because substances appeared to have different degrees of each of these building blocks. For example, a burning, green stick releases smoke (air), water, and ash (earth), and since it burns, it obviously contains fire. The Aristotelian view lasted for almost 2,000 years.

Modern Theories of Matter.

Dalton's Atomic Theory. Dalton proposed atomic theory in 1803 to explain his observations about the relative masses of elements in a compound. Dalton's Postulates say: [1] Matter is made of small particles called atoms; [2] Atoms are indestructible. They cannot be created or destroyed during chemical or physical changes; [3] Atoms of an element are identical. They have the same mass; [4] Atoms of different elements have different masses; [5] Compounds are formed by combining atoms of different elements. The Dalton model of the atom is a solid, indivisible sphere.

Thomson's Model. Thomson showed that the beam of light in a Crookes tube was actually composed of negatively charged particles he called electrons. In order to account for the negatively charged particles in neutral matter, he assumed the rest of the atom was positively charged. Thomson's model of the atom was a positively-charged cloud with the negative electrons scattered evenly throughout.



Rutherford's Model. Ernest Rutherford performed the alpha scattering experiment in 1911. He probed the inside of the atom by aiming small, positively charged particles called alpha particles at gold foil. Only 1 in 8,000 alpha particles bounced straight back or were deflected greatly. The rest went straight through the gold foil. Based on these observations, Rutherford suggested that the atom is mostly empty space with a small, positively charged center and negatively charged electrons revolving around the outside like planets around the sun.

The Bohr Model. Bohr developed a model of the atom with circular pathways for the electron. These pathways were at fixed distances from the nucleus. Electrons could be found only in these circular pathways. If an electron absorbed enough energy, it could jump up to another level, but it could never be found between levels. Inevitably, the electron lost energy and fell back down to a previous level, giving off the extra energy as a specific frequency of light. Bohr had a complex equation into which he could substitute simple numbers, integers such as 1, 2, or 3, and the equation predicted the frequencies of the light. These mystery numbers represented the energy levels of the electrons. Bohr's evidence for his model was the bright line spectra formed when electrons fell from the excited state back to the ground state.

Answer the questions below by circling the number of the correct response.

- In 1661, Robert Boyle proposed that matter is composed of simple substances called elements that cannot be broken down by chemical means. His theory was not accepted for over 100 years because: (1) he lacked qualitative evidence, (2) he lacked quantitative evidence, (3) he lacked both qualitative and quantitative evidence, (4) he was a philosopher rather than a scientist.
- Boyle proposed that matter is composed of simple substances called elements that cannot be further decomposed or broken down. Aristotle's idea that air was an element did not fit Boyle's definition because air: (1) cannot be seen, (2) cannot be broken down, (3) can be decomposed, (4) can react with other substances.
- Antoine Lavoisier measured the mass of some mercury. Then he heated it in air. The mercury turned into a red substance as it was heated. The mass of the red substance was greater than the mass of the mercury. The best explanation for this observation is that: (1) matter was created, (2) a new element formed, (3) the mercury combined with something in the air, (4) hot mercury is heavier than cold mercury.
- Lavoisier heated mercury until a red substance formed. Then he heated the red substance and it broke down into mercury and oxygen. From his observations, it is reasonable to conclude: (1) mercury and oxygen are elements and the red substance is a compound composed of mercury and oxygen, (2) mercury, oxygen, and the red substance are elements, (3) the red substance is an element, and mercury and oxygen are compounds composed from it, (4) mercury, oxygen, and the red substance are compounds.

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