

- Q6 When water was added to the anhydrate of $\text{CuSO}_4(s)$, did the reaction proceed in the direction of the reactants or products?

B.2 Reaction of Copper(II) Ion and Hydroxide

	Test tube 1	Test tube 2
Initial color of CuCl_2 solution (1)	lighter blue liquid transparent	lighter blue liquid transparent
Initial appearance $\text{Cu}(\text{OH})_2(s)$ (2)	cloudy blue solution	cloudy blue solution
Change in appearance (3)	more cloudy solution	
Change in appearance (4)		clearer lighter blue solution

- Q7 In test tube 1, did the addition of OH^- cause the reaction to go in the direction of the reactants or products? Explain.

- Q8 In test tube 2, did the addition of H_3O^+ cause the reaction to go in the direction of the reactants or products? Explain.

C. Changing Equilibrium Conditions: Le Châtelier's Principle

C.1 Effect of Changing Reactant Concentration

Test tube	Color	Is color deeper or lighter red compared to the reference?	$[\text{FeSCN}^{2+}]$ increases or decreases?
1. Reference	blood red		
2. Fe^{3+} added	dark blood red	deeper	
3. SCN^- added	dark blood red	deeper	
4. Fe^{3+} removed	light blood red	lighter	

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**Barron's Educational Series, Neil D.
Jespersen, Pamela Kerrigan**



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