

ANA EXPRESS Plotter Elite Series

Applicable Models: AE-60E, AE-75E, AE-101E, AE-120E

REPLACEMENT PARTS PRICE LIST

<i>PART NO.</i>	<i>DESCRIPTION</i>	<i>AE-60E</i>	<i>AE-75E</i>	<i>AE-101E</i>	<i>AE-120E</i>	<i>SUGGESTED LIST PRICE</i>
18000-28	Blade Holder Assembly	X	X	X	X	124.00
18042-01	Blade Holder Cap	X	X	X	X	16.00
17066-01	Blade Holder Dowel Pin	X	X	X	X	4.00
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19009-02	Bottom Pan		X			132.00
19009-03	Bottom Pan			X		168.00
19009-04	Bottom Pan				X	198.00
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19075-04	Bottom Pan Assembly		X			1150.00
19101-02	Bottom Pan Assembly			X		1250.00
19120-02	Bottom Pan Assembly				X	1350.00
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19060-01	Clutch O-Ring, Roller		X	X	X	2.00
19000-28	Clutch, Assembly, Vinyl Roller (New Type)	X	X	X	X	40.00
19000-12	Flex Cable (to Carriage Head)	X				46.00
19075-11	Flex Cable (to Carriage Head)		X			58.00
19101-04	Flex Cable (to Carriage Head)			X		74.00
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19021-02	Head Cover	X	X	X	X	10.00
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18078-01	Idler Pulley Bearing	X	X	X	X	10.00
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19050-01	Nut, Lock, Self Threading, 2-56	X	X	X	X	1.00
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18000-19	Paper Sensor Assembly	X				24.00
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web elements in the second row of the periodic table place their electrons in the $2n$ shell as well as the $1n$ shell for instance lithium li has three electrons two fill the $1s$ orbital and the third is placed in the $2s$ orbital giving an electron configuration of $1s^2 2s^1$

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web one more point needs to be emphasized about the relationship between electron configuration and the periodic table the atoms of elements in the same vertical column of the table have similar electron configurations for example consider the alkaline earth elements group iia

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are $he\ 2$ helium $1s^2$ $ne\ 10$ neon $1s^2\ 2s^2\ 2p^6$

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web q1 what is the electron configuration of the element potassium answer the atomic number of potassium is 19 a neutral potassium element has 19 electrons its electron configuration is $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^1$ or simply put $ar\ 4s^1$ q2 what is the total number of electrons that can be filled in s p and d subshells

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blocks depending on their electron configuration