

Periodic Trends Worksheet

Directions: Use your class materials to complete the following questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, O, S, Cl

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

K, Na, S, O, Cl

3. Which atom has the higher first ionization energy: sodium or sulfur?

Sulfur (sulfur is further from the nucleus than sodium and has a higher effective nuclear charge)

4. Which two elements in the same family generally have similar properties?

There have been two cases each where a third element with similar properties exists (boron, aluminum, gallium, indium)

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (decreasing to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest?

Seventh period, francium (francium is further from the nucleus than any other element)

7. Which period in the periodic table shows atoms with the smallest atomic radius? Which element has the smallest?

Second period, fluorine (fluorine is closest to the nucleus than any other element)

8. Write the atoms in each group that have the largest atomic radius.

Li	Na	K
Rb	Cs	Fr
Be	Mg	Ca
Sc	Y	La
Ti	Zr	Hf
V	Nb	Ta
Cr	Mo	Tc
Mn	Co	Ni
Fe	Cu	Zn
Ag	Au	Hg

Al	B	P
Sb	As	Bi
Sn	Ge	Se
Pb	Si	Te
Bi	P	As
At	Te	S
Cl	Br	I
At	At	At

9. Circle the atoms in each group that have the greatest electronegativity.

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Wolfgang Guggemos

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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