

LOTS OF  
CONVERSIONS!Name Grove  
Period      Date     /    /    **8 • Why Do Hot Air Balloons Float?****THE IDEAL GAS LAW**

$$PV = nRT \text{ where}$$

P = pressure in atmosphere

V = volume in liters

n = number of moles of gas

R = Universal Gas Constant = 0.0821 L·atm/mol·K

T = Kelvin temperature

See  
scratch  
paper

- How many moles of oxygen will occupy a volume of 2.50 liters at 1.20 atm and ~~25°C~~ <sup>298 K</sup>?  

$$PV = nRT \quad n = \frac{PV}{RT} = \frac{1.20 \text{ atm} \cdot 2.50 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 298 \text{ K}} = 0.947 \text{ mol}$$
- What volume will 2.00 moles of nitrogen occupy at ~~720 mmHg~~ <sup>0.947 atm</sup> and ~~20°C~~ <sup>293 K</sup>?  

$$V = \frac{nRT}{P} = \frac{2.00 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 293 \text{ K}}{0.947 \text{ atm}} = 500 \text{ L}$$
- What pressure will be exerted by ~~25.0 g~~ <sup>0.568 mol</sup> of  $\text{CO}_2$  at temperature of ~~25°C~~ <sup>298 K</sup> and a volume of ~~500 mL~~ <sup>0.500 L</sup>?  

$$P = \frac{nRT}{V} = \frac{0.568 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 298 \text{ K}}{0.500 \text{ L}} = 0.947 \text{ atm}$$
- At what temperature will ~~5.00 g~~ <sup>0.704 mol</sup> of  $\text{Cl}_2$  exert a pressure of ~~900 mmHg~~ <sup>1.18 atm</sup> at a volume of ~~250 mL~~ <sup>0.250 L</sup>?  

$$T = \frac{PV}{nR} = \frac{1.18 \text{ atm} \cdot 0.250 \text{ L}}{0.704 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}} = 293 \text{ K}$$
- What is the ~~density~~ <sup>9/L</sup> of  $\text{NH}_3$  at ~~800 mmHg~~ <sup>1.05 atm</sup> and ~~25°C~~ <sup>298 K</sup>?  

$$V = 1.00 \text{ L} \quad n = ?$$
- If the density of a gas is 1.2 g/L at ~~745 mmHg~~ <sup>0.980 atm</sup> and ~~20°C~~ <sup>293 K</sup>, what is its molar mass?  $\frac{2}{\text{mol}} = 1.20 \text{ g/mol}$
- How many moles of nitrogen gas will occupy a volume of ~~347 mL~~ <sup>0.347 L</sup> at ~~6680 mmHg~~ <sup>8.79 atm</sup> and ~~27°C~~ <sup>300 K</sup>?  

$$n = \frac{PV}{RT} = \frac{8.79 \text{ atm} \cdot 0.347 \text{ L}}{0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 300 \text{ K}} = 0.120 \text{ mol}$$
- What volume will ~~454 grams~~ <sup>225 mol</sup> (1 lb) of hydrogen occupy at 1.05 atm and ~~25°C~~ <sup>298 K</sup>?  

$$V = \frac{nRT}{P} = \frac{225 \text{ mol} \cdot 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K} \cdot 298 \text{ K}}{1.05 \text{ atm}} = 500 \text{ L}$$
- Find the number of grams of  $\text{CO}_2$  that exert a pressure of ~~285 mmHg~~ <sup>1.03 atm</sup> at a volume of 32.5 L and a temperature of ~~32°C~~ <sup>305 K</sup>.  

$$n = ?$$
- An elemental gas has a mass of 10.3 g. If the volume is 58.4 L and the pressure is ~~758 mmHg~~ <sup>0.997 atm</sup> at a temperature of ~~25°C~~ <sup>275.5 K</sup>, what is the gas? Find molar mass

# Ideal Gas Law Chemistry If8766 Answers

**R Bogdan**



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web bathsheba s role in these few terse lines reveals very little of her feeling or character she plays an almost entirely passive role in this chapter and utters only three words two in hebrew the parenthetical aside about her purifying herself may refer back to the roof bath as a ritual cleansing at the end of her period which would affirm

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template 1 for poor performance date employer s full name office address re warning notice for poor performance dear employee s name your performance is consistently falling short of our company standards you aren t meeting assigned deadlines delivering quality work or reaching your individual goals

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jan 12 2021    send and email or a written request for a meeting use the company letterhead to issue the warning serve the letter in a formal setting and never in a public space or at his or her desk clearly state your company policy regarding the

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