

### **Problem 1**

Using the Ideal Gas Law, answer the following questions:

- a) How would the pressure of the gas change if the number of molecules of the gas were doubled as well as doubling the volume, without changing the temperature of the gas?
  
  
  
  
  
  
  
  
  
  
- b) The temperature of the gas is held fix. If the pressure is reduced to one-fourth its initial value and the volume is expanded to twice its value, how would the number of molecules of the gas change?
  
  
  
  
  
  
  
  
  
  
- c) The temperature (in K) of the gas were doubled, while the pressure decreases to one-eighth its initial value and the number of molecules is held fix. How would the volume change from its initial value?

# Ideal Gas Law Answers

**Kathleen Armour**

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