

সময়—৩ ঘণ্টা

পূর্ণমান—৬০ + ৪০ = ১০০

সৃজনশীল প্রশ্ন

মান—৬০

দ্রষ্টব্য :—জান পাশের সংখ্যা প্রশ্নের পূর্ণমান জ্ঞাপক। প্রতিটি বিভাগ থেকে কমপক্ষে একটি করে প্রশ্ন নিয়ে মোট ছয়টি প্রশ্নের উত্তর দাও।

ক বিভাগ—পাঠ্যগণিত

- | | | |
|-----|--|-------|
| ১। | একটি আয়তাকার বাগানের দৈর্ঘ্য ৪০ মিটার এবং প্রস্থ ৩০ মিটার। বাগানের ভেতর চারিদিকে ৩ মিটার চওড়া একটি রাস্তা আছে। ২৫ সে.মি. দৈর্ঘ্য এবং ১২.৫ সে. মি. প্রস্থবিশিষ্ট ইট দিয়ে রাস্তাটি পাকা করা হলো। | নম্বর |
| (ক) | বাগানের ক্ষেত্রফল নির্ণয় কর। | ২ |
| (খ) | রাস্তার ক্ষেত্রফল নির্ণয় কর। | ৪ |
| (গ) | রাস্তাটি পাকাকরণে প্রয়োজনীয় ইটের সংখ্যা বের কর। | ৪ |
| ২। | বাবুল হোসেন ৮% ক্ষতিতে একটি ছাগল বিক্রয় করলেন। কিন্তু ছাগলটি ৮০০ টাকা বেশি মূল্যে বিক্রয় করলে তাঁর ৮% লাভ হত। পরবর্তীতে তিনি ছাগলটির ক্রয়মূল্যের সমপরিমাণ টাকা বার্ষিক ১০% মুনাফায় ৩ বছরের জন্যে ব্যাংকে জমা রাখলেন। | |
| (ক) | ৮% সরল মুনাফায় ৮০০ টাকার ৩ বছরের মুনাফা নির্ণয় কর। | ২ |
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| (গ) | উক্ত টাকা ব্যাংকে জমা রাখলে সরল মুনাফা ও চক্রবৃদ্ধি মুনাফার পার্থক্য কত হবে? | ৪ |

খ বিভাগ—বীজগণিত

- | | | |
|-----|--|---|
| ৩। | $U = \{x : x \text{ স্বাভাবিক সংখ্যা এবং } x < 8\};$
$P = \{1, 3, 5\};$
$Q = \{x : x \text{ স্বাভাবিক সংখ্যা এবং } 1 < x \leq 7\}$ এবং
$R = \{2, 4, 6\}.$ | |
| (ক) | $P \cap Q$ নির্ণয় কর। | ২ |
| (খ) | $(P \cap Q) \cup (Q \cup R)$ নির্ণয় কর। | ২ |
| (গ) | প্রমাণ কর যে, $(Q \cup R)' = Q' \cap R'.$ | ৪ |
| ৪। | $3x + 2y = 12, 2x + 3y = 13$ দুটি সমীকরণ— | |
| (ক) | $(4, 0)$ বিন্দুটি প্রথম সমীকরণকে সিদ্ধ করে কিনা যাচাই কর। | ২ |
| (খ) | অপনয়ন পদ্ধতিতে সমীকরণদ্বয়ের সমাধান কর। | ৪ |
| (গ) | লেখচিত্র থেকে সমীকরণদ্বয়ের সমাধান নির্ণয় কর। | ৪ |
| ৫। | $x + \frac{1}{x} = 5$ একটি বীজগণিতিক সমীকরণ যেখানে $x > 0.$ | |
| (ক) | $3x + \frac{3}{x}$ এর মান নির্ণয় কর। | ২ |
| (খ) | $x^4 + \frac{1}{x^4}$ এর মান নির্ণয় কর। | ৪ |
| (গ) | প্রমাণ কর যে, $x^2 - \frac{1}{x^2} \neq x^3 + \frac{1}{x^3}$ | ৪ |

গ বিভাগ—জ্যামিতি

- | | | |
|-----|--|---|
| ৬। | $ABCD$ একটি আয়তক্ষেত্র যার $AB = 3$ সে. মি. উহার AC ও BD কর্ণদ্বয় পরস্পরকে O বিন্দুতে ছেদ করে। | |
| (ক) | উন্নীপকের তথ্যগুলো চিত্রের সাহায্যে প্রকাশ কর। | ২ |
| (খ) | প্রমাণ কর যে, $AC = BD$ এবং $AO = CO, BO = DO.$ | ৪ |
| (গ) | AB কে একটি বাহুর দৈর্ঘ্য এবং $AC = 5$ সে. মি.-কে একটি কর্ণের দৈর্ঘ্য ধরে একটি রহস্য অঙ্কন কর। | ৪ |

বাকী পৃষ্ঠাগুলি নিচের লিঙ্কে ক্লিক করে ডাউনলোড করে পড়ুন।

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KJ Lindholm-Leary



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web ideal gas law google classroom a sample of h a 2 g is contained in a cylinder with a moveable piston at an initial pressure of p 1 the volume of the sample is increased from 3 0 l to 6 0 l while the temperature is held constant as shown in the diagram below

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web answer as temperature of a gas increases pressure will also increase based on the ideal gas law the volume of the tire can only expand so much before the rubber gives and releases the build up of pressure

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web the observed behavior of gases embodied in the empirical gas laws leads to a series of equations that can be summarized by a single equation of state called the ideal gas law equation this shows the relationship between a gas s pressure p temperature t volume v and amount in moles n

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web each gas in a mixture of gases exerts its own pressure independently of the other gases present therefore the pressure of each gas within a mixture is called the partial pressure of the gas dalton s law of partial pressures states that the total pressure of a mixture of gases is equal to the sum of the partial pressures of the component gases

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web this unit explores the physical nature of gases the laws governing the behavior of gases and applications of gases from air bags to ozone depletion the gas laws and the ideal gas equation

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web the following practice problems are to master to topics on the ideal gas laws boyle s law charles s law and avogadro s law as well as the combined gas law equation there are examples to work on the dalton law of partial pressures the graham s law of effusion and gas stoichiometry

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web at very high pressures a gas will occupy a larger volume than predicted by the ideal gas law which of the following statements is true at low temperatures intermolecular forces become important and the pressure of a gas will be lower than predicted by