

Intermolecular Forces Quiz

Part 1: Complete this on this document.

A 1. Which molecule will have the strongest IMFs:

a. ammonia (NH_3)

b. chlorine gas (Cl_2)

c. methane (CH_4)

d. CH_3F

D 2. Which one of the following substances has London dispersion force as its only intermolecular force?

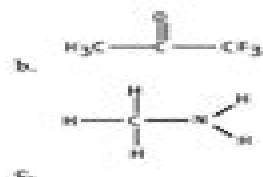
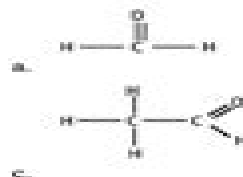
a. CH_3OH

b. NH_3

c. H_2S

d. CF_4

D 3. Which one of the following substances would have hydrogen bonding as one of its intermolecular forces (assume it will be surrounded by other molecules exactly like itself)?



A 4. Hydrogen bonding is a covalent bond.

a. always true

b. sometimes true

C c. never true

C 5. The weaker the IMFs are in a liquid substance, the _____ (the) drops of liquid can be placed on a penny without it falling off, and the time for it to completely evaporate will be _____.

a. more, longer

b. more, shorter

c. fewer, longer

D d. fewer, shorter

B 6. Sugar is added to a beaker of pure water. The structure for sucrose, a type of sugar, is seen below. Predict the strongest IMF between sugar molecules.

a. ion-dipole

b. hydrogen bonding

c. dipole-dipole

d. Dispersion



Short answer:

5. Explain in detail why/how washing your hands with soap and water (instead of just water) is effective at killing bacteria, germs, and the coronavirus. In your explanation, you should use many chemistry terms. More is better here.

The soap has a hydrophobic end and a hydrophilic end. The soap molecules become micelles which are soluble in water. The hydrophilic side is suspended in water and the hydrophobic side traps the germs so they can't come in contact with water.

Intermolecular Forces Test Questions High School

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