



## **Assessment Guide Grade 5**



# **Assessment Guide**



# Grade 5 Assessment Guide

**David Baud**



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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the

amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$  problems 6 & 8 are similar to problem 5

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web limiting reagents problem one 1 for the balanced equation  $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$  if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  $\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g/mol}} = 0.666 \text{ moles } \text{O}_2$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$   $\frac{160.0 \text{ g } \text{Al}}{27 \text{ g/mol}} \times \frac{1 \text{ mol } \text{Al}_2\text{O}_3}{2 \text{ mol } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 302.3 \text{ g } \text{Al}_2\text{O}_3$  c determine the percent yield  $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h<sub>2</sub> are left over 13

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