7.5 × 10<sup>-4</sup>. What is the value of equilibrium constant [RPMT-2001] (2) 1/3(4) None (1)3(3)1Q.21If the equilibrium constant for the reaction [RPMT-2002] N<sub>2</sub> + 3H<sub>2</sub> → 2NH<sub>3</sub>, is K then the equilibrium, constant for the reaction  $2N_1 + 6H_2 \rightarrow 4NH_3$ , would be equal to  $(4) \frac{1}{122}$ (3) (2) JK (1)  $\mathbb{K}^{n}$ The reaction quotient (Q) for the reaction [CBSE-PMT-2003]  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$  is given by  $Q = \frac{[NH_3]^2}{[N_2][H_2]^3}$ . The reaction will proceed from right to left if (1)  $Q < K_s$  (2)  $Q > K_o$  (3) Q = 0[CBSE-PMT-2003] The following equilibria are given Q.23  $N_2 + 3H_2 \rightarrow 2NH_3 \dots K_1$   $N_3 + O_2 \rightarrow 2NO \dots K_2$  $H_1 + \frac{1}{2}O_2 \rightarrow H_2O$  ......  $K_1$ The equilibrium constant of the reaction  $2NH_3 + \frac{5}{2}O_2 \rightarrow 2NO + 3H_2O$ , in terms of  $K_1 \cdot K_2$  and  $K_3$ 15 (1) K<sub>1</sub> K<sub>2</sub> / K<sub>3</sub> (2) K<sub>1</sub> K<sub>2</sub> / K<sub>3</sub> (3) K<sub>2</sub> K<sub>3</sub> / K<sub>1</sub> (4) K<sub>1</sub> K<sub>2</sub> K<sub>3</sub> For the gas phase reaction, 0.24 $C_1H_4 + H_2 \longrightarrow C_2H_4$  ( $\Delta H = -32.7$  kcal). carried out in a vessel, the equilibrium conc. of C,H, can be increased by [AMU 2003] Decreasing the pressure (2) Increasing the temperature (3) Removing some C,H (4) Adding some H, Consider the following reversible reaction at equilibrium,  $2H_{*}O(g) \Longrightarrow 2H_{*}(g) + O_{*}(g)$ Q.25  $\Delta H = 24.7$  kJ. Which one of the following changes in conditions will lead to maximum decomposition of H<sub>2</sub>O(g)? [Kerala CEE 2004] (1) Increasing both temperature and pressure (2) Decreasing temperature and increasing pressure (3) Increasing temperature and decreasing pressure (4) Increasing temperature at constant pressure In the reaction,  $A_{s}(g) + 4B_{s}(g) \rightleftharpoons 2AB_{s}(g) \Delta H < 0$ , the formation of  $AB_{s}(g)$  will be favoured at Q.26[MP PMT 2004] (1) High temperature, low pressure (2) Low temperature, high pressure (3) High temperature, high pressure (4) Low temperature, low pressure

For a reversible reaction forward and backward rate constant respectively are 2.5 × 10<sup>-4</sup> and

Q.20

# **Mcq On Environmetal Chemistry**

**Priti Singhal** 

# Mcq On Environmetal Chemistry:

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