

11 Which compound has the greatest percent composition by mass of sulfur?

- (1) BaS (2) CaS (3) MgS (4) SrS

way #64

12 Two molecules of HBr collide and then form H_2 and Br_2 . During the collision, the bonds in the HBr molecules are

- (1) broken as energy is absorbed
(2) broken as energy is released
(3) formed as energy is absorbed
(4) formed as energy is released

way #43

BA/RF

13 Which atom in the ground state has a stable electron configuration?

- (1) carbon (2) magnesium (3) neon (4) oxygen

way #49

14 Which statement describes a multiple covalent bond?

- (1) Two electrons are shared.
(2) Four electrons are shared.
(3) Two electrons are transferred.
(4) Four electrons are transferred.

way #46

15 The electronegativity difference between the atoms in a molecule of HCl can be used to determine

- (1) the entropy of the atoms
(2) the atomic number of the atoms
(3) the first ionization energy of the atoms
(4) the polarity of the bond between the two atoms

way #51

16 Which two gases can not be broken down by chemical means?

- (1) CO and He (2) CO and NH_3 (3) Xe and He (4) Xe and NH_3

way #5

elements Table S

17 Two substances in a mixture differ in density and particle size. These properties can be used to

- (1) separate the substances
(2) chemically combine the substances
(3) determine the freezing point of the mixture
(4) predict the electrical conductivity of the mixture

18 Which unit is used to express an amount of thermal energy?

- (1) gram (2) mole (3) joule (4) pascal

Table D

19 Under which conditions of temperature and pressure does a real gas behave most like an ideal gas?

- (1) low temperature and low pressure
(2) low temperature and high pressure
(3) high temperature and low pressure
(4) high temperature and high pressure

way #9

20 According to the kinetic molecular theory for an ideal gas, all gas particles

- (1) are in random, constant, straight-line motion
(2) are separated by very small distances relative to their sizes
(3) have strong intermolecular forces
(4) have collisions that decrease the total energy of the system

way #8

21 Which mathematical expression represents the heat of reaction for a chemical reaction?

- (1) (the heat of fusion) - (the heat of vaporization)
(2) (the heat of vaporization) - (the heat of fusion)
(3) (the potential energy of the products) - (the potential energy of the reactants)
(4) (the potential energy of the reactants) - (the potential energy of the products)

way #78

22 At 101.3 kPa and 298 K, a 1.0-mole sample of which compound absorbs the greatest amount of heat as the entire sample dissolves in water?

- (1) LiBr -48.83 (2) NaCl +2.38 (3) NaOH -46.51 (4) NH_4Cl +14.8

way #11

Table I

23 For a reaction at equilibrium, which change can increase the rates of the forward and reverse reactions?

- (1) a decrease in the concentration of the reactants
(2) a decrease in the surface area of the products
(3) an increase in the temperature of the system
(4) an increase in the activation energy of the forward reaction

June 2003 Physics Regents Answers Explained

Miriam Lazar



June 2003 Physics Regents Answers Explained:

Testing Regimes, Accountabilities and Education Policy Bob Lingard, Goli Rezai-Rashti, Wayne Martino, 2017-10-02

Around the globe various kinds of testing including high stakes national census testing have become meta policies steering educational systems in particular directions and having great effects on schools and on teacher practices as well as upon student learning and curricula There has also been a complementary global aspect to this with the OECD s PISA and IEA s TIMSS and PIRLS which have had impacts on national education systems and their policy frameworks While there has been a globalized educational policy discourse that suggests that high stakes standardised testing will drive up standards and enhance the quality of a nation s human capital and thus their international economic competitiveness this discourse still manifests itself in specific vernacular path dependent ways in different nations High stakes testing and its effects can also be seen as part of the phenomenon of the datafication of the world and policy as numbers linked to other reforms of the state including new public management network governance and top down and test based modes of accountability This edited collection provides theoretically and empirically informed analyses of these developments This book was originally published as a special issue of the Journal of Education Policy

Twenty-first Century Schools Gerard MacDonald, David Hursh, 2006-01-01 Contemporary school systems are not working well Educational solutions abound but the problems remain This is because our school systems are not primarily concerned with education Their driving forces are political and economic Any systemic solution to schooling s current difficulties will start with politics not in any conventional sense with education Twenty first Century Schools traces the extension of political control over Britain s school system and through US case studies looks at alternative methods of organisation The authors argue that Anglo American school systems provide a good education for a small minority and to the majority offer inadequate schooling Though this has always been inequitable it may in the past have been economically efficient at least from the viewpoint of the state If that was once true for mercantile and industrial economies it is not true now The knowledge economies toward which the UK and US are moving demand a continuing ability to learn and to innovate right across the workforce Our time honoured tradition of anachronistic curricula didactically taught does not develop these qualities For those reasons the current Anglo American drive to entrench unequal educational opportunity is at once economically politically and socially misconceived From these premises Twenty first Century Schools goes on to outline the political and educational changes needed to shape school systems which are socially and economically adapted to the new century

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