

Chemistry of Life

Study Guide B**Answer Key****SECTION 1. ATOMS, IONS, AND MOLECULES**

1. An element is a certain type of atom.
2. Shultz should remember the use of the abbreviations in Figure 1.2 in Section 1. Protons should be in the center with neutrons and electrons. Elements should be labeled in writing surrounding the nucleus.
3. A compound is composed of atoms of different elements.
4. an atom that has gained or lost one or more electrons
5. the number of protons in the nucleus equals the number of electrons
6. Shultz should remember the first abbreviations in Figure 1.3, with + for acids, losing an electron (becoming a positive ion) and the electron accepting that electron (becoming a negative ion).
7. A chemical bond is between two atoms that share a pair of electrons.
8. the number of electron pairs that is needed to share for all the outer energy level
9. molecule
10. compound
11. element
12. ion
13. Ionic bonds form between oppositely-charged ions. covalent bonds form when atoms share a pair of electrons.

SECTION 2. PROPERTIES OF WATER

1. a molecule with a slightly positively-charged oxygen and a slightly negatively-charged hydrogen
2. The oxygen nucleus pulls the electrons in the molecule more strongly than the hydrogen atom. The hydrogen atom has a slight positive charge and the oxygen atom has a slight negative charge.
3. an attraction between a slightly positive hydrogen and a slightly negative atom
4. between the slightly positive hydrogen atom of one water molecule and the

slightly negative oxygen atom of another water molecule

5. A large amount of energy is needed to produce an increase in water temperature that results changes in temperature.
6. "Water molecule" "mole" for each other
7. "Water molecule" "mole" for other substances
8. A solvent is present in greater concentrations and dissolves the solute.
9. increased pump efficiency
10. smaller molecules, such as the solute
11. **Effect on H⁺ concentration:** acids increase H⁺ concentration; bases decrease H⁺ concentration
Effect on pH: acids lower pH; bases raise pH
12. Shultz should remember the first abbreviations in the first section. Definition: homeostasis: maintenance of substances within a substance present in greater amount in which other substances allowing water molecules that therefore is a solvent.

SECTION 3. CARBOHYDRATE MOLECULES

1. Carbon atoms join the bases of the molecules that make up most living things.
2. Carbon atoms can form covalent bonds with other carbon atoms, including when carbon atoms.
3. Shultz should remember Figure 3.1 in Section 3.
4. carbon atoms are a source of chemical energy part of cell structure
5. sugars, starches, cellulose
6. carbon atoms are a source of chemical energy part of cell structure
7. fats, oils, phospholipids
8. many functions including movement, transport, chemical catalyst
9. vacuoles, homeostasis
10. store genetic information, build proteins

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Kathleen Armour



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1. Understanding the eBook Holt Mcdougal Biology Study Guide B Chemistry Of Life
 - The Rise of Digital Reading Holt Mcdougal Biology Study Guide B Chemistry Of Life
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Mcdougal Biology Study Guide B Chemistry Of Life
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Mcdougal Biology Study Guide B Chemistry Of Life
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Mcdougal Biology Study Guide B Chemistry Of Life
 - Personalized Recommendations
 - Holt Mcdougal Biology Study Guide B Chemistry Of Life User Reviews and Ratings
 - Holt Mcdougal Biology Study Guide B Chemistry Of Life and Bestseller Lists
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 - Holt Mcdougal Biology Study Guide B Chemistry Of Life eBook Subscription Services
 - Holt Mcdougal Biology Study Guide B Chemistry Of Life Budget-Friendly Options
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- ePub, PDF, MOBI, and More
 - Holt Mcdougal Biology Study Guide B Chemistry Of Life Compatibility with Devices
 - Holt Mcdougal Biology Study Guide B Chemistry Of Life Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Mcdougal Biology Study Guide B Chemistry Of Life
 - Highlighting and Note-Taking Holt Mcdougal Biology Study Guide B Chemistry Of Life
 - Interactive Elements Holt Mcdougal Biology Study Guide B Chemistry Of Life
 8. Staying Engaged with Holt Mcdougal Biology Study Guide B Chemistry Of Life
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Mcdougal Biology Study Guide B Chemistry Of Life
 9. Balancing eBooks and Physical Books Holt Mcdougal Biology Study Guide B Chemistry Of Life
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Mcdougal Biology Study Guide B Chemistry Of Life
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Holt Mcdougal Biology Study Guide B Chemistry Of Life
 - Setting Reading Goals Holt Mcdougal Biology Study Guide B Chemistry Of Life
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Holt Mcdougal Biology Study Guide B Chemistry Of Life
 - Fact-Checking eBook Content of Holt Mcdougal Biology Study Guide B Chemistry Of Life
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
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