

Half Modern Chemistry Review

CHAPTER 5: THE PERIODIC TABLE

The following pages contain the **best** (but not all) of the information for this chapter is that.

Focus on this content, but make sure to review other topics: reactions, bonding, molecules, etc.

If you study this document and **UNDERSTAND** what you should at least be able to **DO** after that.

**** That does not mean revisit, reanalyze, reapply or redefine of this content. ****

OUTCOMES:

- Communicate with regard to subatomic/chemical/atomic structure. (P 42)
- Communicate chemical content to quantify and predict. (P 42)
- 5.1: Describe the organization of the modern Periodic table: distinguish between periods and groups. (P 42)
- 5.2: Use the Periodic Table to obtain information about the properties (physical and chemical, corrosion, energy, and atomic size) of that elements. (P 42) (P43)
- 5.3: Identify common compounds, metalloids, and noble gases; identify the representative groups and/or main groups (groups 1, 2, 13, 14, 15, 16, 17, 18) (P 42 & 43)
- 5.4: Compare periodic trends of atomic size, ionization energy, and electronegativity. (P 42 & 43)

5.1: HISTORY OF THE PERIODIC TABLE

• Introductory

- **periodic law** = the law that states that the repeating chemical and physical properties of elements change periodically with the atomic numbers of the elements.
- **periodic table** = an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column.
- **classification** = a method of the arrangement of elements, based on their chemical, physical, and other properties. (P 42)
- **metalloids** = one of the series of heavy metalloids elements that include from element number 44 through element 100 (element 100 is the periodic table).

• Chapter Highlights

- The periodic law states that the physical and chemical properties of the elements, and periodic functions of their atomic numbers.
- The periodic table is an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column/group/family.
- The columns in the periodic table are referred to as groups.

5.2: ELECTRON CONFIGURATIONS AND THE PERIODIC TABLE

• Introductory

- **atomic shell** = one of the elements of Group 1 of the periodic table (lithium, sodium, potassium, rubidium, cesium, and francium).
- **subatomic shell** = one of the elements of Group 2 of the periodic table (beryllium, magnesium, calcium, strontium, and barium).
- **representative elements** = one of the metals that are not the inner shell metals using the outer shell to form.
- **metalloids** = one of the elements of the periodic table (boron, silicon, germanium, arsenic, antimony, and tellurium).
- **halogens** = one of the elements of Group 17 (fluorine, chlorine, bromine, iodine, and astatine).
- **halogens** = one of the elements of Group 18 (helium, neon, argon, krypton, xenon, and radon).

Holt Modern Chemistry Test Guide Answers

Marco Cascella



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